

Nanostructured Ceria-Based Materials: Effect of the Hydrothermal Synthesis Conditions on the Structural Properties and Catalytic Activity

*Original*

Nanostructured Ceria-Based Materials: Effect of the Hydrothermal Synthesis Conditions on the Structural Properties and Catalytic Activity / Piumetti, Marco; Bensaid, Samir; Andana, Tahrizi; Dosa, Melodj; Novara, Chiara; Giorgis, Fabrizio; Russo, Nunzio; Fino, Debora. - In: CATALYSTS. - ISSN 2073-4344. - ELETTRONICO. - 7:6(2017), pp. 1-18. [10.3390/catal7060174]

*Availability:*

This version is available at: 11583/2676571 since: 2017-07-17T19:05:22Z

*Publisher:*

MDPI AG

*Published*

DOI:10.3390/catal7060174

*Terms of use:*

This article is made available under terms and conditions as specified in the corresponding bibliographic description in the repository

*Publisher copyright*

(Article begins on next page)

Article

# Nanostructured Ceria-Based Materials: Effect of the Hydrothermal Synthesis Conditions on the Structural Properties and Catalytic Activity

Marco Piumetti, Samir Bensaid \*, Tahrizi Andana, Melodj Dosa, Chiara Novara, Fabrizio Giorgis, Nunzio Russo and Debora Fino

Department of Applied Science and Technology, Politecnico di Torino, Corso Duca degli Abruzzi 24, 10129 Turin, Italy; marco.piumetti@polito.it (M.P.); tahrizi.andana@polito.it (T.A.); melodj.dosa@polito.it (M.D.); chiara.novara@polito.it (C.N.); fabrizio.giorgis@polito.it (F.G.); nunzio.russo@polito.it (N.R.); debora.fino@polito.it (D.F.)

\* Correspondence: samir.bensaid@polito.it; Tel.: +39-011-090-4662

Academic Editors: Agustín Bueno-López, Enrique Rodríguez-Castellón and Elisa Moretti

Received: 7 April 2017; Accepted: 27 May 2017; Published: 2 June 2017

**Abstract:** In this work, several nanostructured ceria catalysts were prepared by means of a hydrothermal procedure, in which the synthesis conditions (i.e., temperature and pH values) were varied. CeO<sub>2</sub> samples of different shapes and structural properties were obtained, namely cubes, rods, cube and nanorod mixtures, and other polyhedra. The prepared materials were tested using four probe catalytic reactions: CO oxidation, NO oxidation, NO<sub>x</sub>-free soot oxidation, and NO<sub>x</sub>-assisted soot oxidation. The physicochemical properties of the prepared catalysts were studied by means of complementary techniques (i.e., XRD, N<sub>2</sub>-physisorption at −196 °C, CO-TPR (temperature-programmed reduction), field emission scanning electron microscopy (FESEM), micro-Raman spectroscopy). The abundance of defects of the catalysts, measured through in-situ Raman spectroscopy at the typical temperatures of each catalytic process, was correlated to the CO and NO oxidation activity of the prepared catalysts, while the soot oxidation reaction (performed in loose conditions), which was hindered by a poor soot-catalyst contact, was found to be less sensitive to the observed structural defects.

**Keywords:** ceria; nanostructures; hydrothermal; basicity; aging temperature; defect; catalytic oxidations; high temperature Raman spectroscopy

## 1. Introduction

Diesel and gasoline-type engines are important sources of outdoor air pollutants, such as volatile organic compounds (VOCs), NO<sub>x</sub>, and particulate matter (PM) [1–4]. Diesel particulate is suspected of causing several health problems, including lung cancer, and as a result studies on diesel engine exhaust have received a great deal of interest over the last few years [1–8]. As a whole, the soot that forms diesel exhaust particulates can be burnt off above 600 °C, and diesel engine exhaust temperatures fall within the 200 to 500 °C range [9]. Thus, many solid catalysts have been developed to carry out oxidation processes at moderate temperatures. In this scenario, ceria-based catalysts have been used for several catalytic processes, i.e., diesel soot combustion, CO oxidation, and VOC abatement [1,4]. Among the physicochemical properties that favor the reactivity of ceria-based materials, both the redox capacity and the availability of surface oxygen species (oxygen storage capacity, OSC) are probably the most important ones [1]. Moreover, recent studies have shown that nanostructured ceria with (100) and (110) planes are usually more active for oxidation processes than conventional (polycrystalline) ceria nanoparticles (NPs) with the preferred exposure of (111) planes [10]. Thus, over the last few decades,

there has been significant progress in the controlled synthesis of CeO<sub>2</sub>-based nanomaterials. Different shapes and structures of CeO<sub>2</sub>-based catalysts have been synthesized using methods designed by various research groups. In general, CeO<sub>2</sub>-NPs can be mentioned, because of their small (featured) size, which endows them with size- and shape-dependent properties, due to the high surface-to-volume ratio (=high fraction of coordinatively unsaturated sites) and unique electronic features (quantum size effects) [11–16]. Therefore, the surface reactivity of CeO<sub>2</sub>-NPs depends not only on the crystal size, but also on their shape and thus on their exposed planes. Moreover, the structural defects (i.e., anion Frenkel pairs and oxygen vacancies) can favor the mobility of charged species, such as electrons or oxygen anions in the solid, and they can therefore be considered pivotal for the generation of catalytic active centers [17,18].

On the other hand, soot-catalyst contact points also affect the soot oxidation activity, and it is therefore necessary to maximize the interaction between the soot particles and the catalyst surface [19].

In the present work, ceria nanomaterials, with different geometric and textural features (i.e., nanocubes, nanorods, high-surface area CeO<sub>2</sub>), were synthesized to investigate the relationship between the synthesis conditions and their structural features, as well as their catalytic activities towards CO, NO, soot, and NO<sub>x</sub>-assisted soot oxidations. The physicochemical properties of the prepared catalysts were studied by means of complementary techniques (i.e., XRD, N<sub>2</sub>-physisorption at –196 °C, CO-TPR (temperature-programmed reduction), field emission scanning electron microscopy (FESEM), micro-Raman spectroscopy).

## 2. Results

The nanostructured ceria samples were synthesized via a hydrothermal synthesis, using NaOH as the precipitating agent. In the present work, two synthesis parameters, namely basicity and aging temperature, were varied to obtain ceria with different shapes, i.e., nanocubes and nanorods.

### 2.1. Textural Properties

Some of the main physical features of the prepared samples, derived from N<sub>2</sub> physisorption at –196 °C and X-ray diffraction (XRD), are reported in Table 1.

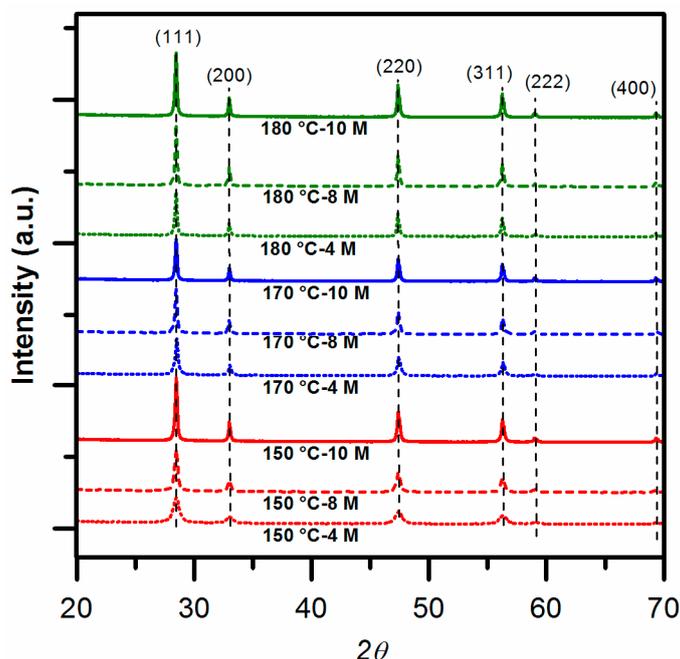
**Table 1.** Textural properties of the samples obtained from N<sub>2</sub> physisorption at –196 °C and X-ray diffraction (XRD) analysis.

Catalyst	S <sub>BET</sub> <sup>1</sup> (m <sup>2</sup> g <sup>–1</sup> )	V <sub>p</sub> <sup>2</sup> (cm <sup>3</sup> g <sup>–1</sup> )	(200)/(111) <sup>3</sup>	(220)/(111) <sup>3</sup>
180 °C-10 M	9	0.02	0.30	0.51
180 °C-8 M	10	0.03	0.30	0.51
180 °C-4 M	14	0.08	0.27	0.52
170 °C-10 M	8	0.02	0.29	0.54
170 °C-8 M	11	0.04	0.28	0.49
170 °C-4 M	28	0.11	0.28	0.53
150 °C-10 M	11	0.03	0.30	0.47
150 °C-8 M	17	0.05	0.26	0.46
150 °C-4 M	52	0.17	0.26	0.51

<sup>1</sup> Specific surface area measured by the Brunauer-Emmett-Teller (BET) method; <sup>2</sup> total pore volume; <sup>3</sup> XRD peak intensity ratio.

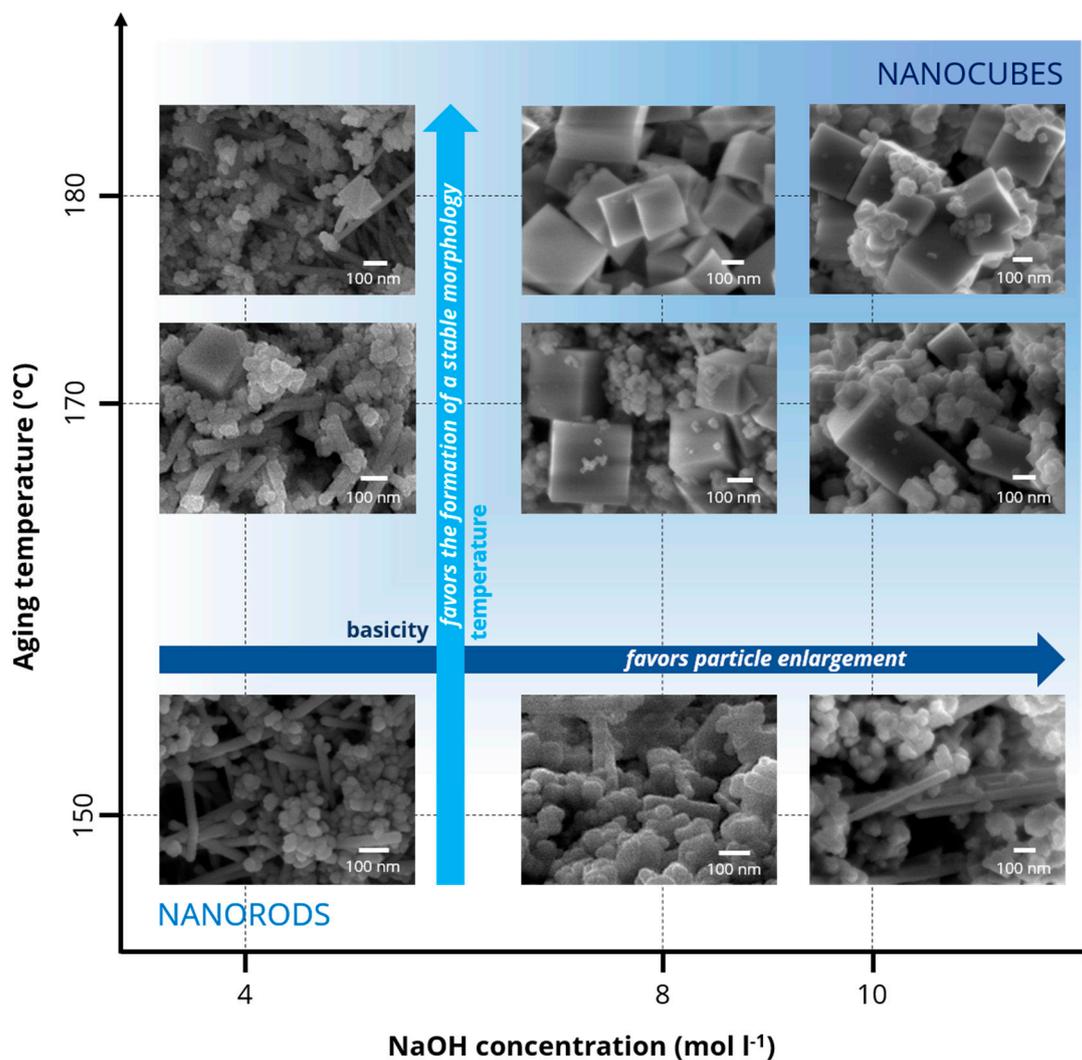
Figure 1 shows the XRD diffractograms of the catalysts. These samples exhibit similar patterns, as far as the cubic fluorite structure is concerned, and are marked by (the existence of) (111), (200), (220), (311), (222), and (400) planes [20,21]. The relative amount of the (110), (100), and (111) surfaces was estimated from the XRD peak intensity ratios. As reported in Table 1, the (200)/(111) ratios have a tendency to increase as a function of the temperature (from 150 to 180 °C) and of the basicity values (NaOH concentration: from 4 to 10 M) adopted during the synthesis; this trend does not appear for the (220)/(111) ratios. On the other hand, the specific surface area (S<sub>BET</sub>) and the total pore volume (V<sub>p</sub>)

of the samples both decrease as the temperature and the NaOH concentration values grow. In other words, it appears that the worse textural properties of the catalysts are associated with more abundant (100) planes, unlike the (110) planes.



**Figure 1.** X-ray diffraction (XRD) patterns of the prepared samples.

Figure 2 shows the FESEM micrographies of the samples prepared at different synthesis conditions:  $x$ -axis = aging temperature (150–180 °C);  $y$ -axis = NaOH concentration (4–10 M). As expected, both the aging temperature and the basicity influence the structural and textural properties of the nanostructured  $\text{CeO}_2$ . It has in fact been observed that, the presence of nanorods (length = 300–500 nm) at lower NaOH concentrations (i.e., 4 M) prevails over the formation of nanocubes (length = 30–50 nm), regardless of the temperature used during the aging step. However, higher temperatures (i.e., 180 °C) seem to favor the creation of larger rods along with a few well-defined nanocubes (side  $\approx$  100 nm). Conversely, the higher basicity favors the development of well-defined cubes when the temperature ranges from 170 to 180 °C. Moreover, both the temperature and the basicity lead to larger and more stable  $\text{CeO}_2$  cubes, with sizes of up to 200–300 nm. On the other hand, a low aging temperature (i.e., 150 °C) leads to a mixture of polyhedra with different shapes and sizes (rods, truncated cubes, etc.). Finally, it should be pointed out that the hydrothermal conditions used for catalyst preparation are not very selective, in terms of morphology, as a multimodal distribution of the cube sizes even appears for samples synthesized at 10 M NaOH and 180 °C.



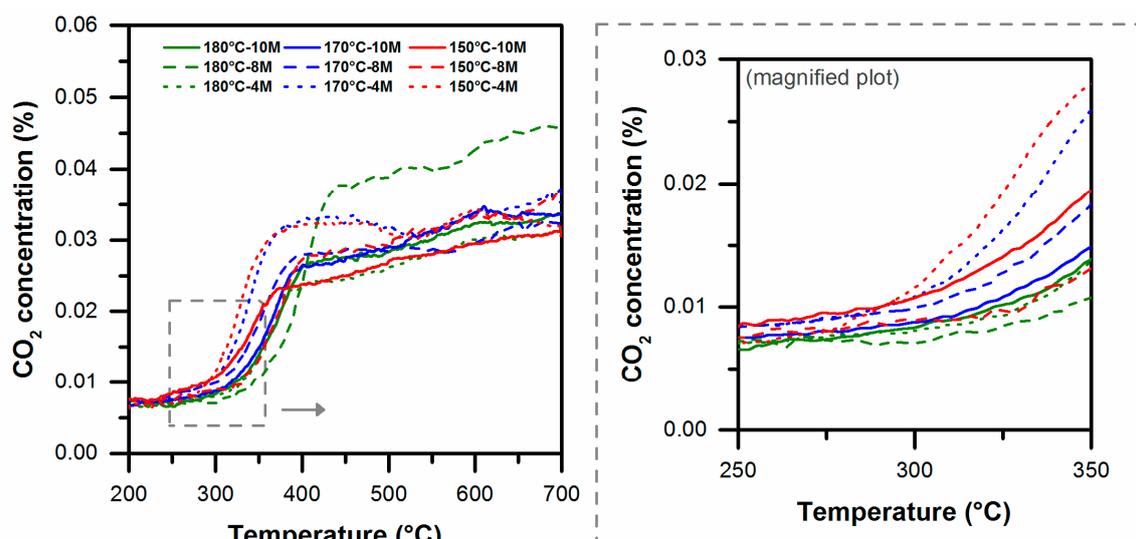
**Figure 2.** Field emission scanning electron microscopy (FESEM) images of the prepared samples as a function of the synthesis conditions (namely, aging temperature and NaOH concentration).

## 2.2. Reducibility and Defective Sites

The reducibility of the catalysts was investigated by means of CO-TPR analysis. CO was selected as the reducing agent, instead of H<sub>2</sub>, due to its abundant presence in diesel exhaust. During the analysis, CO reacts with the oxygen atoms that originate solely from the solid catalyst, and CO<sub>2</sub> is formed as a product. The nature of the oxygen species on the catalyst surface can be understood by considering the reduction peak temperature, which is estimated via deconvolution, since the oxygen release is highly temperature-dependent. As a whole, a good catalyst for CO and soot oxidation should be able to release oxygen at a temperature that is compatible with the expected field of application.

Figure 3 shows the CO-TPR curves of the prepared samples, all of which seem to conform to the usual trend for pure ceria: an early emergence of CO<sub>2</sub> at around 300–400 °C and a steady production of CO<sub>2</sub> up to the end of the test. The early formation of CO<sub>2</sub> signifies the consumption of labile oxygen species on the catalyst surface. Constriction points can be observed above 500 °C in some samples, and these indicate a change in the oxygen source for the reaction, which possibly comes from the bulk phase of the sample. Table 2 summarizes the peak temperatures and their corresponding estimated quantities of surface oxygen (mmol) obtained from the curve deconvolution. The earliest catalyst reduction was demonstrated for the 170 °C-4 M and 150 °C-4 M samples (reduction peak temperatures at 388 °C and 373 °C, respectively). The samples synthesized with high aging temperatures and basicity, namely,

the 180 °C-10 M and 180 °C-8 M samples, exhibit lower reducibility (reduction peak temperatures at 399 °C and 424 °C, respectively). The results suggest that CO reduction on ceria catalysts depends to a great extent on the textural properties. The high surface area of the ceria nanorods synthesized at low aging temperature leads to more exposure of the catalyst surface to CO, thus enabling early CO<sub>2</sub> production. Despite the lower reducibility, the 180 °C-8 M sample appears to release a much higher quantity of oxygen (about 0.029 mmol). It has been surmised that, while the surface area determines the degree of reducibility, the structure of the catalyst is responsible for the amount of oxygen that can be exploited for the reaction.



**Figure 3.** CO<sub>2</sub> production as a function of temperature over all the samples derived from CO-TPR (temperature-programmed reduction) analysis.

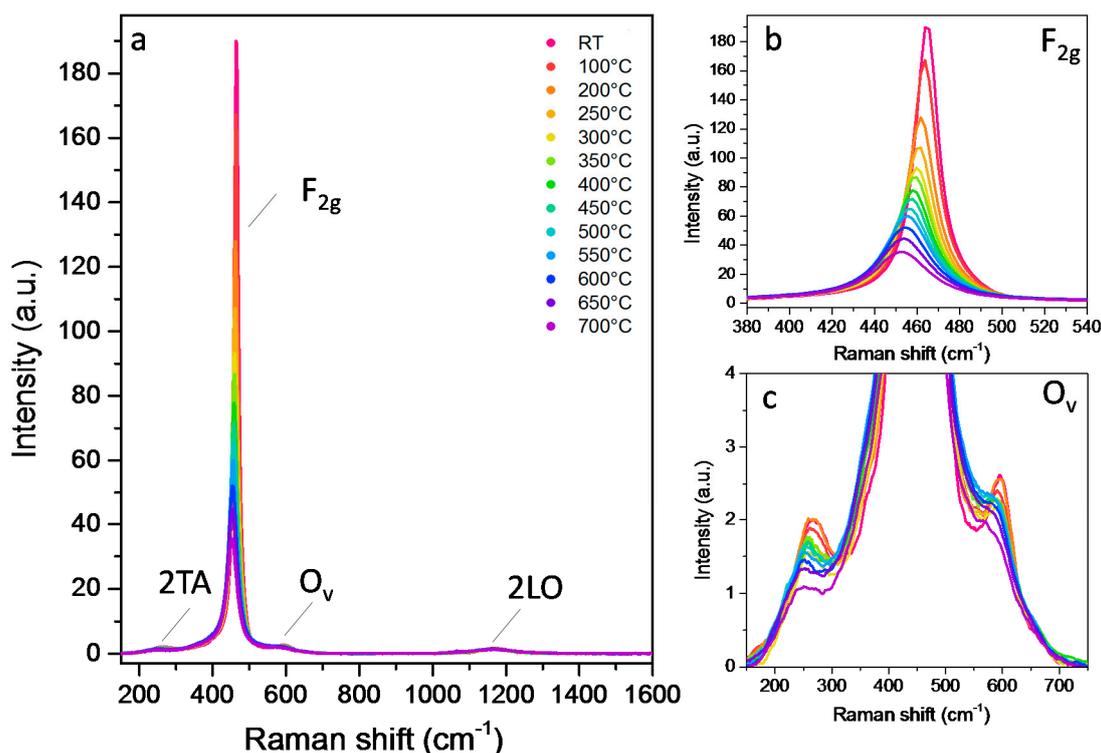
**Table 2.** Estimated quantity of surface oxygen and specific reaction rate (calculated at 300 °C) obtained from the CO-TPR (temperature-programmed reduction) curve deconvolution.

Catalyst	Estimated Quantity of Surface Oxygen (Mmol)	Specific Reaction Rate (mmol <sub>CO</sub> g cat <sup>-1</sup> h <sup>-1</sup> )
180 °C-10 M	0.023	164
180 °C-8 M	0.029	42.6
180 °C-4 M	0.013	142
170 °C-10 M	0.021	207
170 °C-8 M	0.025	309
170 °C-4 M	0.024	402
150 °C-10 M	0.018	393
150 °C-8 M	0.019	220
150 °C-4 M	0.022	473

The samples were further analysed by means of micro-Raman spectroscopy to study how the synthesis parameters affect their structural features and thus their catalytic activity. Raman scattering is a powerful tool that can be used to directly probe the defect sites in doped and undoped ceria particles [22–24]. In order to improve the reliability of the characterization, in-situ Raman measurements were performed at several temperatures, starting from room temperature (r.t.) to 700 °C, in order to cover the range in which catalytic phenomena take place.

Figure 4a reports the typical behavior of the vibrational modes of a representative ceria NP sample (180 °C-8 M) at different temperatures. Four features can be identified in the room temperature spectra: the main Ce-O<sub>8</sub> crystal unit vibration around 464 cm<sup>-1</sup>, which is the characteristic F<sub>2g</sub> mode of the

fluorite lattice [25], and three weaker Raman peaks at  $261\text{ cm}^{-1}$ ,  $598\text{ cm}^{-1}$ , and  $1165\text{ cm}^{-1}$ . The first and the third ones can be assigned to the higher order modes of ceria, namely to the second order transverse acoustic mode, that is, 2TA [24] (although a different assignment to a Ce-OH related vibration has recently been suggested [26]), and to the second overtone of the longitudinal optical band (2LO) [26]. The Raman peak at  $598\text{ cm}^{-1}$  is instead a disorder induced mode arising from Frenkel-type anion defects, consisting of a vacancy, originated by the motion of an  $\text{O}^{2-}$  ion to an interstitial octahedral site [17,23]. Defect related bands around  $550\text{ cm}^{-1}$ , due to oxygen vacancies at a  $\text{Ce}^{3+}$  site [23], cannot be resolved, thus pointing out that this set of ceria NPs mainly contains oxidized  $\text{Ce}^{4+}$ .



**Figure 4.** (a) Raman spectra of the 180 °C-10 M sample acquired from room temperature to 700 °C. (b) Enlarged  $\text{F}_{2g}$  band region. (c) Enlarged  $\text{O}_v$  band region.

Upon increase in temperature, the most significant Raman spectral changes are observed for the peak position, linewidth, and intensity of the  $\text{F}_{2g}$  mode. An enlarged view of the temperature dependent spectra in the  $\text{F}_{2g}$  region is shown in Figure 4b. A red-shift of the  $\text{F}_{2g}$  mode occurs, together with an asymmetric broadening of its line shape and an intensity decrease, when the samples are heated. This phenomenon has been observed for several materials [27–29] as well as for ceria nanocrystals [30,31], and has been explained considering two different contributions: thermal expansion and lattice anharmonicity [29]. While the former only influences the peak position, both Raman shift and linewidth variations are consequences of the phonons anharmonic decay at higher temperatures, which may involve three- or four-phonon interactions [31].

Figure 4c provides a zoomed view of the region of the defect peaks. Similarly to the  $\text{F}_{2g}$  mode, the defect band at  $598\text{ cm}^{-1}$  is subjected to a red shift and to a relevant linewidth broadening when the temperature is increased.

According to the existing literature, the defect density in the catalyst can be estimated on the basis of the intensity ratio between the  $\text{O}_v$  and  $\text{F}_{2g}$  bands,  $\text{O}_v/\text{F}_{2g}$  (the cumulative Raman band areas ascribed to the different defect sites are usually considered) [22,32].

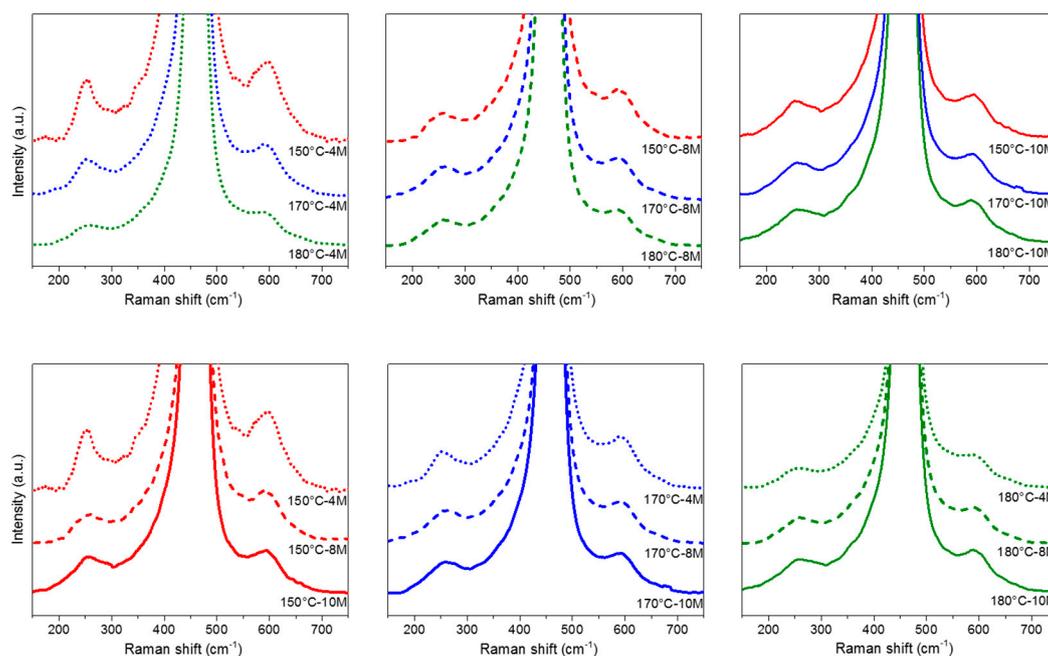
Such a parameter is particularly useful, as it has been shown to correlate well with the catalytic performances of the analyzed materials [32,33]. Nevertheless, the room temperature Raman spectra

of the catalysts are usually analyzed, while more reliable information could be provided by high temperature Raman measurements. Table 3 reports the calculated  $O_v/F_{2g}$  ratio at 300 °C, 400 °C, and 500 °C. Because of the intensity decrease in the  $F_{2g}$  and defect bands during heating, the ratios can only be compared at a fixed temperature. The influence of anharmonicity on a collective vibrational mode, such as the  $F_{2g}$  one, is in fact stronger than that observed for localized modes, such as defect induced vibrations, and as a result, the ratio generally increases at higher temperatures.

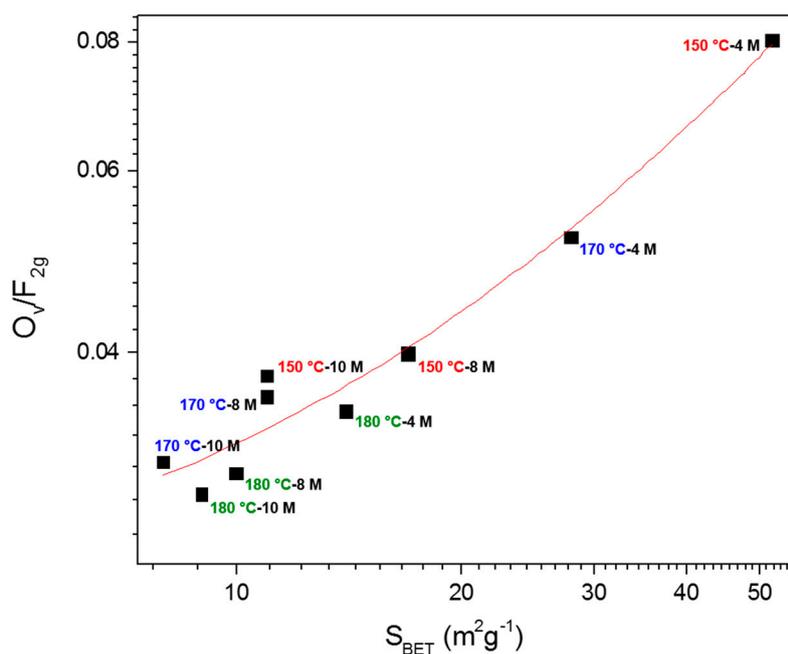
**Table 3.** Intensity ratios between the  $F_{2g}$  and  $O_v$  bands derived from micro-Raman spectroscopy. Spectra recorded at 300, 400, and 500 °C.

Catalysts	$O_v/F_{2g}$ 300 °C	$O_v/F_{2g}$ 400 °C	$O_v/F_{2g}$ 500 °C
150 °C-10 M	0.038	0.041	0.039
170 °C-10 M	0.031	0.035	0.036
180 °C-10 M	0.029	0.035	0.040
150 °C-8 M	0.040	0.044	0.048
170 °C-8 M	0.036	0.037	0.037
180 °C-8 M	0.031	0.031	0.032
150 °C-4 M	0.080	0.093	0.100
170 °C-4 M	0.052	0.052	0.063
180 °C-4 M	0.035	0.036	0.038

Figure 5 shows the Raman spectra of the catalysts collected at 300 °C. The best discrimination of the catalytic performances towards the CO oxidation is in fact expected around that temperature window. The spectra are normalized to the  $F_{2g}$  peak and grouped according to the main synthesis parameters: NaOH concentration (top) and aging temperature (bottom). It can be inferred from the figure, which is the graphical equivalent of the  $O_v/F_{2g}$  ratio, that an increase in the NaOH concentration or aging temperature tends to reduce the number of defects, and the influence of the synthesis temperature is greater. Additionally, as shown in Table 3, when specimens with different NaOH concentrations are compared, the highest variability is observed for the samples synthesized at 150 °C, while the ratios are closer for the NPs prepared at 180 °C. Figure 5 also highlights that highly defective surfaces are characterized by the broadest  $F_{2g}$  linewidth. These results can clearly be correlated to the morphological and textural features of the NPs. Actually, smaller nanoparticles usually show larger  $F_{2g}$  linewidths, due to the combination of phonon confinement and strain effects [24]. Accordingly, the linewidth of the  $F_{2g}$  band decreases moving from lower (small rods and cubes) to higher aging temperatures, the latter of which produce larger nanoparticles. Moreover, the comparative analysis of the FESEM images and of the Raman spectra reveals that as the NPs dimension increases, the defect density measured by the  $O_v/F_{2g}$  ratio is reduced. Such behavior was previously observed for samples of different sizes, but with the same composition and morphology [24,32]. In the present study, a direct correlation has been found between the  $O_v/F_{2g}$  ratio and the specific surface area of the  $CeO_2$  NPs (Figure 6), as the latter parameter decreases when the size of the NPs increases. In contrast to previous works, the defect density measured by Raman spectroscopy is weakly influenced by the exposed crystal planes, while the dependence on the specific surface area is dominant.



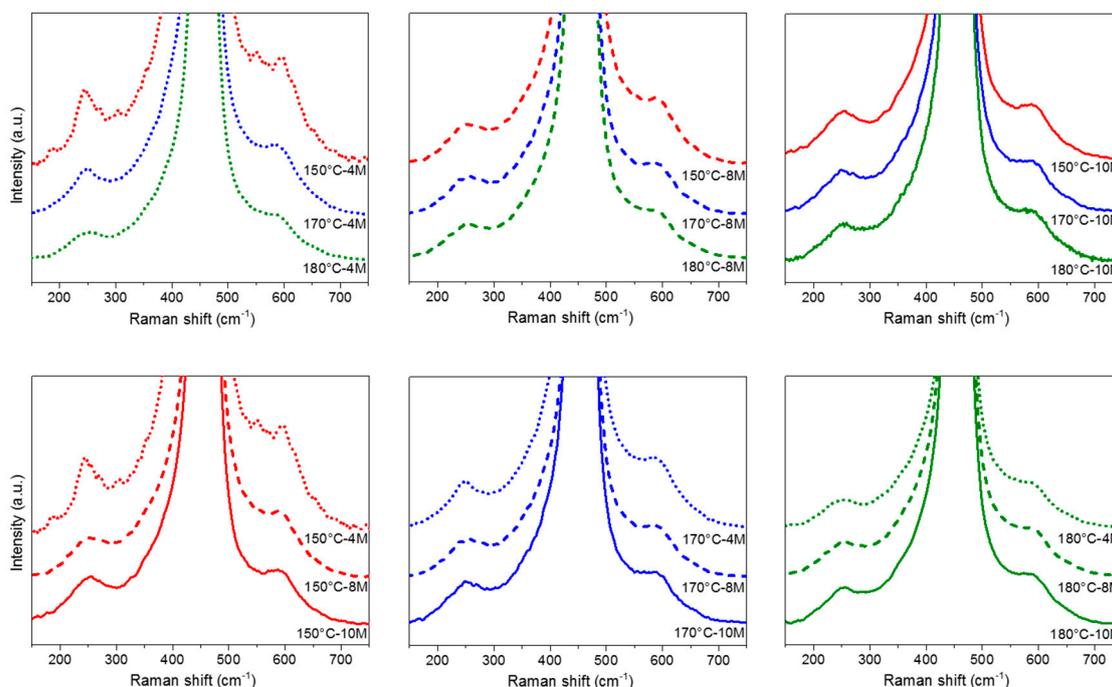
**Figure 5.** Raman spectra of the CeO<sub>2</sub> samples at 300 °C. The spectra are grouped according to the synthesis conditions: NaOH concentration (**top**); aging temperature (**bottom**).



**Figure 6.** Logarithmic plot of the  $O_V/F_{2g}$  ratio versus the specific surface area of each CeO<sub>2</sub> sample, showing a good correlation between the two parameters. The red line is a linear fit of the experimental data.

Higher temperatures should be considered for the soot conversion reaction. For this reason, the normalized spectra at 500 °C are reported in Figure 7. Although the defect bands are significantly broader and differences between the samples are partially mitigated, the  $O_V/F_{2g}$  trend is analogous to the one detected for the 300 °C measurements. It should be noted that no evident shoulder can be detected at 550 cm<sup>-1</sup>, due to the formation of oxygen vacancies associated to the reduction of Ce<sup>4+</sup>, during heating. Moreover, after the thermal cycle from r.t. to 700 °C, a reversibility of the

spectral features can be observed for all the CeO<sub>2</sub> NPs, not only for the F<sub>2g</sub> peak, as already reported for CeO<sub>2</sub> nanoparticles with a size >20 nm [24], but also for the defect band. This suggests that the spectral changes revealed at high temperatures are due more to temperature effects on the vibrational properties of the samples than to an evolution of the microstructure of the samples. The observed reversibility is also in agreement with the reported stability of Frenkel-type anion defects, for which the onset of recombination phenomena was shown to be over 700 °C [34].



**Figure 7.** Raman spectra of the CeO<sub>2</sub> samples at 500 °C. The spectra are grouped according to the synthesis conditions: NaOH concentration (**top**); aging temperature (**bottom**).

### 2.3. Catalytic Activity

The catalytic activity for CO oxidation was studied in the 75–500 °C temperature range. Figure 8 shows the CO conversion to CO<sub>2</sub> as a function of the reaction temperature of the prepared catalysts along with the uncatalyzed reaction. All the catalysts exhibit positive conversion trends for increasing temperature, and total oxidation of CO is reached at about 250 °C for the 150 °C-4 M sample. This sample in fact exhibited the best results in terms of T<sub>10%-50%-90%</sub> values (197, 218, and 231 °C, respectively) (Table 4), thus confirming the beneficial role of the textural properties on the overall activity of this set of CeO<sub>2</sub> nanostructures. In other words, a higher surface area means more abundant surface defects (i.e., oxygen vacancies) which in turn are accessible to the reactants. According to the literature [35], CO molecules reach the Ce<sup>n+</sup>-O species and form CO<sub>2</sub>, and then the O vacancy is readily refilled by subsurface (lattice) oxygen. As expected, poor textural properties lead to lower catalytic performances (i.e., in terms of CO<sub>2</sub> conversion), and the least performing catalyst was found to be the ceria 180 °C-10 M, which exhibits low S<sub>BET</sub> and V<sub>p</sub>. These catalytic results are moderately correlated to the CO-TPR profiles, thus suggesting that the presence of molecular oxygen plays a role in this reaction.

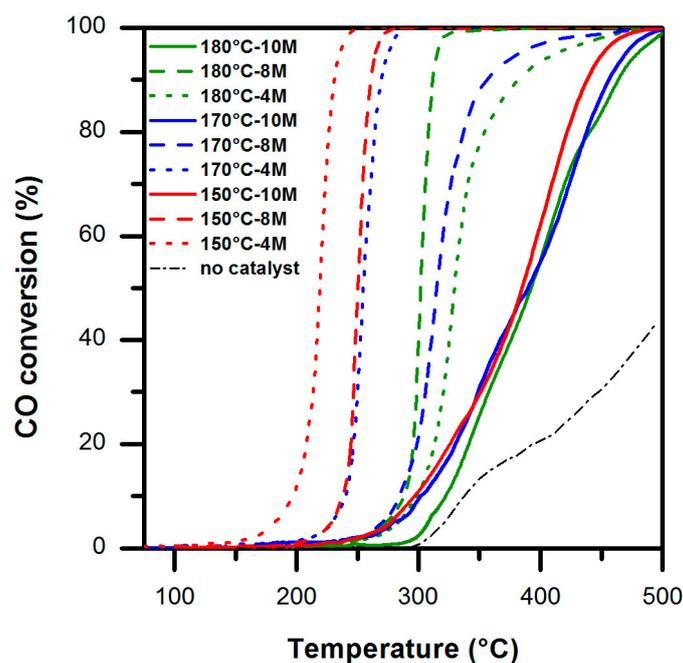


Figure 8. CO to CO<sub>2</sub> conversion (%) versus temperature over the prepared catalysts.

Table 4. Results from the soot oxidation tests in the absence and (the) presence of NO<sub>x</sub>.

In the Absence of NO <sub>x</sub> (10%-v O <sub>2</sub> in N <sub>2</sub> )			
Catalyst	T <sub>conversion</sub> (°C)		
	T <sub>10%</sub>	T <sub>50%</sub>	T <sub>90%</sub>
180 °C-10 M	499	576	618
180 °C-8 M	476	564	614
180 °C-4 M	500	590	639
170 °C-10 M	512	599	640
170 °C-8 M	526	593	631
170 °C-4 M	504	583	629
150 °C-10 M	512	591	637
150 °C-8 M	474	535	618
150 °C-4 M	490	588	635
No catalyst	526	615	656
In the Presence of NO <sub>x</sub> (550 ppmv-NO + 10%-v O <sub>2</sub> in N <sub>2</sub> )			
Catalyst	T <sub>conversion</sub> (°C)		
	T <sub>10%</sub>	T <sub>50%</sub>	T <sub>90%</sub>
180 °C-10 M	480	576	624
180 °C-8 M	442	530	598
180 °C-4 M	468	558	610
170 °C-10 M	471	570	622
170 °C-8 M	469	554	602
170 °C-4 M	461	546	597
150 °C-10 M	468	554	609
150 °C-8 M	412	490	552
150 °C-4 M	440	517	565
No catalyst	526	615	656

Figure 9 summarizes the results derived from the NO oxidation tests, in which the NO<sub>2</sub> percentage in the NO<sub>x</sub> mixture is represented as a function of the temperature. The catalysts synthesized at 180 °C

(green curves) generally demonstrate lower NO oxidation activity. A maximum NO conversion of about 35% was reached over these catalysts at about 475 °C. The catalysts synthesized at 170 °C and 150 °C at a higher basicity (namely, 8 M and 10 M) exhibit a similar conversion trend: the maximum conversion values are centered at  $\approx 475$  °C, with the intensity narrowly varying between 35% and 38%. Surprisingly, the 170 °C-4 M and 150 °C-4 M samples give the highest conversion of NO to NO<sub>2</sub>. The peak of the former sample is centered at 447 °C (conversion  $\approx 40\%$ ), while that of the latter is centered at 421 °C (conversion  $\approx 49\%$ ). The catalytic performance of the 150 °C-4 M sample is in fact rather similar to that of several Pt-based catalysts reported in the literature [4,36]. The high activity of the last two catalysts may be attributed to their relatively high surface area (28 and 52 m<sup>2</sup> g<sup>-1</sup> for 170 °C-4 M and 150 °C-4 M, respectively). The other samples in the set, irrespective of the morphology and particle size, exhibit similar oxidation activities, most likely because they exhibit comparable surface areas (8–14 m<sup>2</sup> g<sup>-1</sup>). Moreover, the NO oxidation trend is similar to that of the CO-TPR profiles in the 200–450 °C range, and a good correlation is thus shown between the NO oxidation reaction and the surface reducibility of ceria.

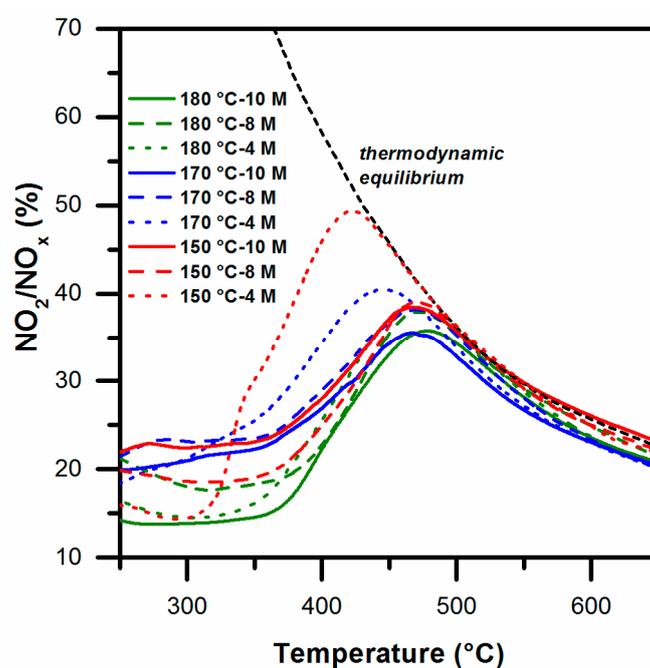


Figure 9. NO oxidation versus temperature over the prepared catalysts.

Figure 10 shows the soot conversion to CO<sub>2</sub> as a function of the temperature obtained with the catalysts in the absence Figure 10 (Sections A–C) and in the presence of NO<sub>x</sub> Figure 10 (Sections D–F) along with Printex-U (no catalyst). Similarly, Table 4 summarizes the data of the temperatures at which 10%, 50%, and 90% soot conversions are reached, extracted from soot conversion curves in the absence and the presence of NO. In the former case, the CeO<sub>2</sub>-containing catalysts exhibit positive soot conversion trends Figure 10 (Section A) for increasing temperature, and soot combustion was completed at about 650 °C for the less reactive materials. On the other hand, uncatalyzed soot combustion takes place at higher temperatures, thus confirming the beneficial role of ceria in this reaction (see Table 4).

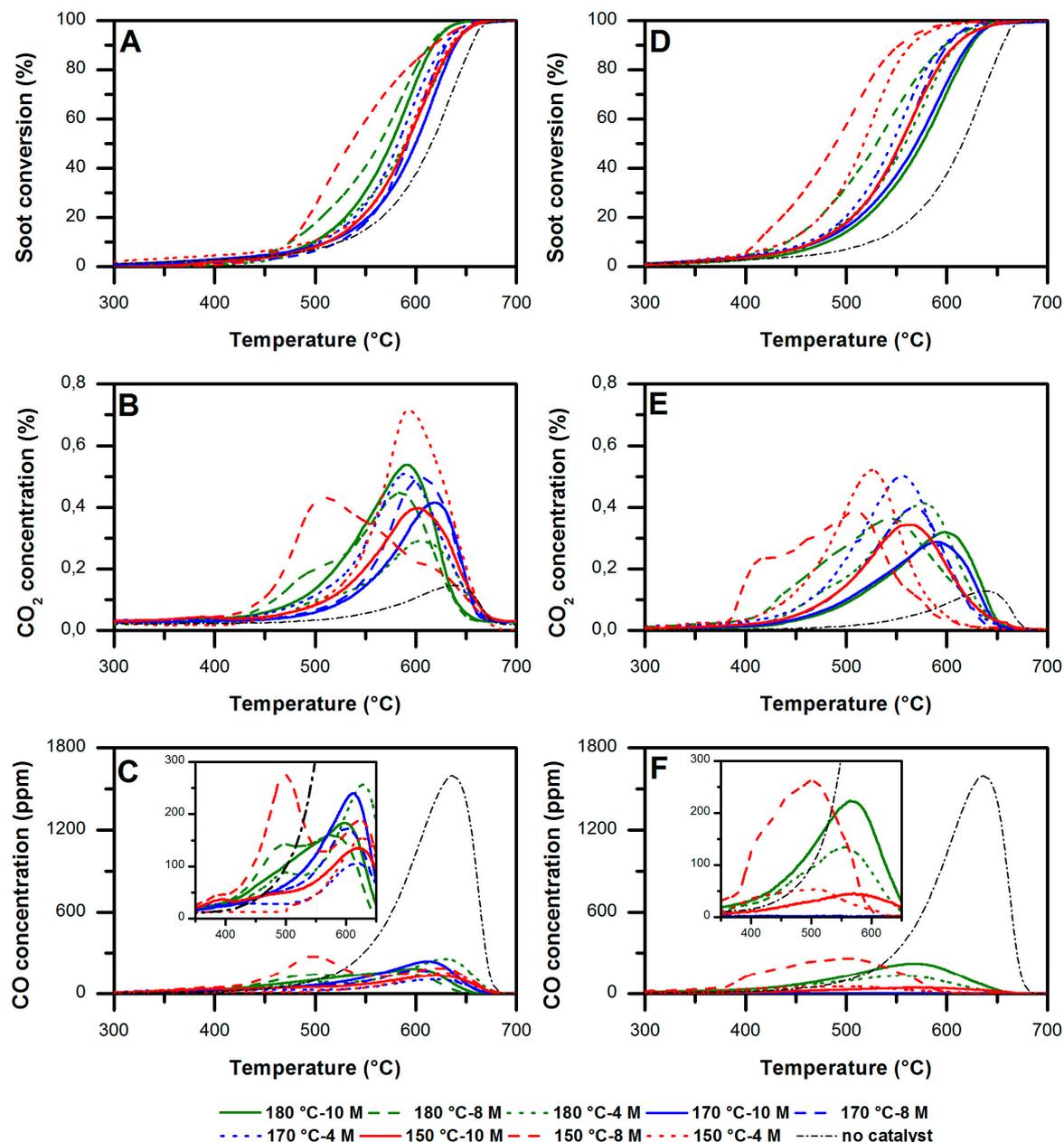


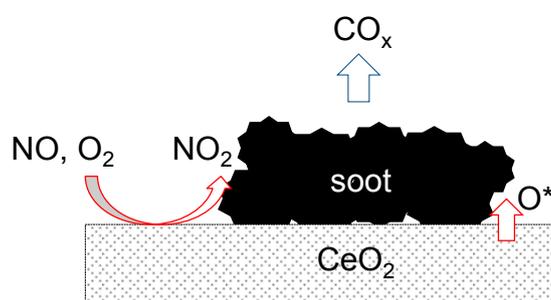
Figure 10. Soot oxidation catalytic tests: (A–C) in the absence; and (D–F) in the presence of  $\text{NO}_x$ .

The 150 °C-8 M sample appears to be the most promising catalyst, and this shows that small polyhedra ( $\approx 50$  nm in size) with abundant (111) planes (see Table 1) can be much more effective than larger cubes (i.e., samples calcined at 170–180 °C, with a basicity of 8–10 M). It has in fact been confirmed in other studies [19], that the soot-catalyst contact conditions play a key role in this solid-solid reaction mediated by gas phase oxygen. On the other hand, the textural properties of the soot oxidation catalysts (i.e., specific surface area, total pore volume) have less effect on the overall oxidation activity.

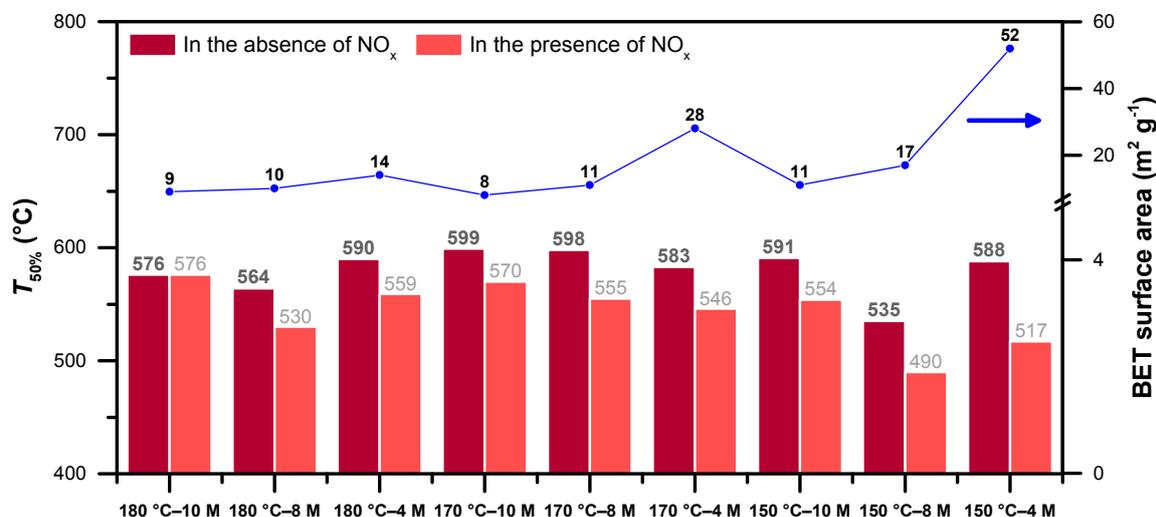
The multimodal distributions of the  $\text{CO}_2$  and CO formations are reported in Sections B and C, respectively. In agreement with the literature [19], the “low-temperature peak” (centered at ca. 500 °C) would seem to reflect the  $\text{CO}_x$  produced by surface redox-type mechanisms, whereas the peak centered at ca. 600 °C can be attributed to the formation of  $\text{CO}_x$  by radical mechanisms that take place at higher temperatures (an almost uncatalyzed reaction). As expected, an intense low-temperature peak appears

for the most effective catalyst (peak centered at ca. 500 °C) along with a shoulder at ca. 610 °C, due to the production of  $\text{CO}_x$  via radical (almost non-catalytic) processes.

Figure 10 (Sections D–F) and Table 4 summarize the results obtained from the soot oxidation tests in the presence of NO. Interestingly, all the catalysts exhibit better performances (in terms of  $T_{10\%}$ ,  $T_{50\%}$  and  $T_{90\%}$ ) for NO-assisted soot combustion than in the absence of NO. The co-presence of  $\text{NO}_x$  and soot on the ceria surface promotes the oxidation activity via multiple (synergistic) effects. Ceria acts as a storage means of  $\text{NO}_x$  under the form of nitrate at moderate temperatures (below 400 °C). Above these temperatures, NO oxidation would be limited, due to thermodynamic equilibrium, while the stored  $\text{NO}_2$  can be released from the ceria to the soot oxidation site [9]. As shown in Scheme 1,  $\text{NO}_x$ -assisted and active oxygen mechanisms ( $\text{O}^*$  derived from spillover phenomena, nitrate, and lattice oxygens) may occur simultaneously during the soot/ $\text{NO}/\text{O}_2$  oxidation reaction. The beneficial role of  $\text{NO}_x$  on soot combustion over ceria appears greater for catalysts with a higher specific surface area, as shown in Figure 11.

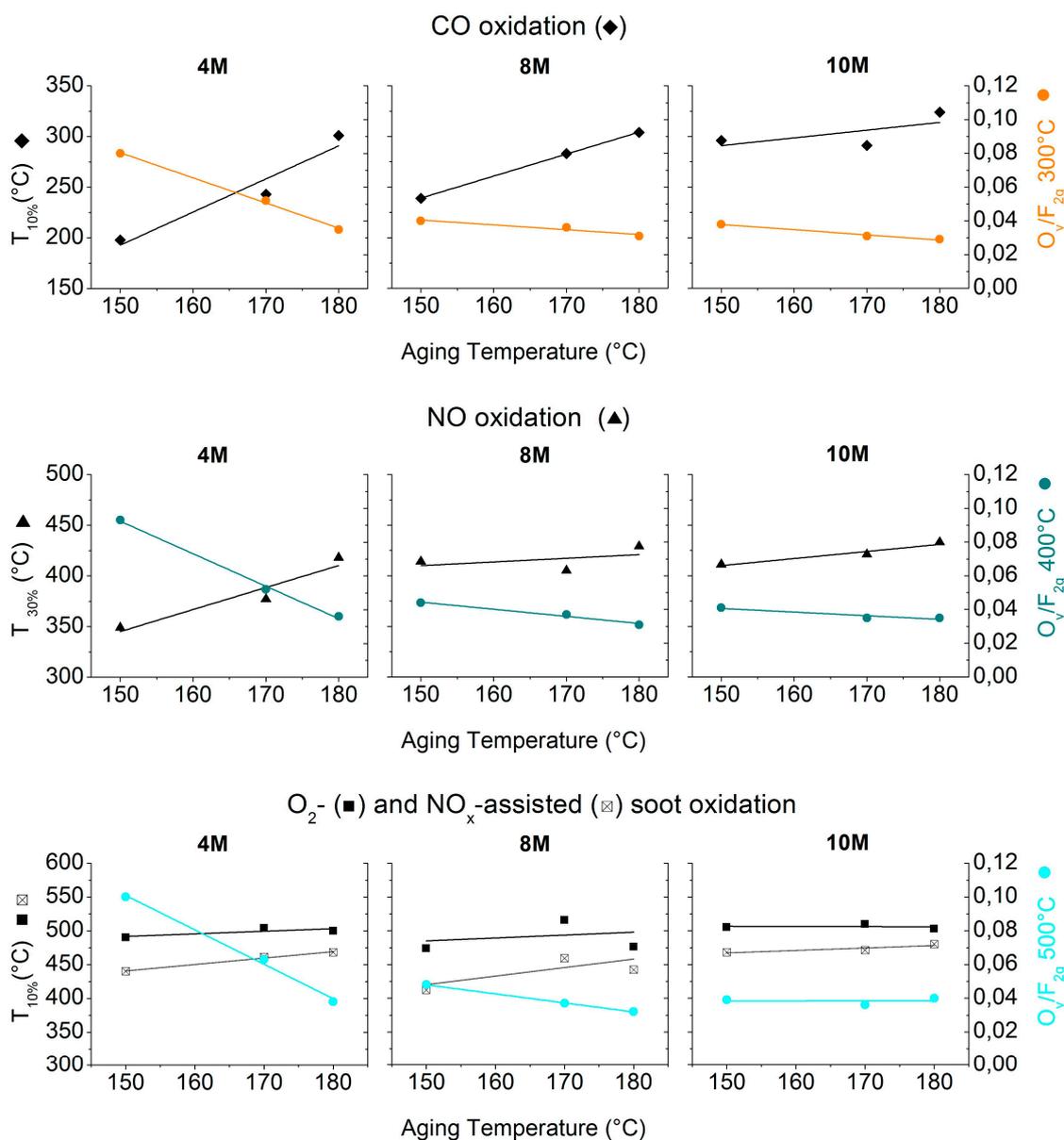


**Scheme 1.** Schematization of the ceria-catalyzed soot combustion under  $\text{NO}_x/\text{O}_2$ .



**Figure 11.** Summary of  $T_{50\%}$  obtained during the test with  $\text{NO}_x$ -free and  $\text{NO}_x$ -assisted soot oxidations.

Finally, the correlation between the catalyst defects measured through the  $\text{O}_v/\text{F}_{2g}$  ratio in the Raman spectra (presented in Table 3) and the onset temperature of the investigated reactions is reported in Figure 12. The catalysts are grouped according to the molarity of the NaOH used during the synthesis procedure, while the  $x$ -axis reports the aging temperature.



**Figure 12.** CO oxidation:  $T_{10\%}$  ( $y$ -axis on the left) and intensity ratio between the  $F_{2g}$  and  $O_v$  bands at 300 °C ( $y$ -axis on the right) vs. catalyst aging temperature. NO oxidation:  $T_{30\%}$  ( $y$ -axis on the left) and intensity ratio between the  $F_{2g}$  and  $O_v$  bands at 400 °C ( $y$ -axis on the right) vs. catalyst aging temperature.  $O_2$ - and  $NO_x$ -mediated soot oxidation in loose contact conditions:  $T_{10\%}$  ( $y$ -axis on the left) and intensity ratio between the  $F_{2g}$  and  $O_v$  bands at 500 °C ( $y$ -axis on the right) vs. catalyst aging temperature. Each plot regroups catalysts prepared with the same NaOH concentration (4, 8, and 10 M).

The CO oxidation activity, expressed as  $T_{10\%}$ , is compared with the structural defect abundance profile obtained with Raman spectroscopy at 300 °C, which is an average value of the representative temperature window for CO oxidation with these materials. As can be seen, a direct correlation exists between activity and defects: the more defective the catalyst structure is, the higher the activity (and therefore the lower the  $T_{10\%}$ ). This evidence applies to the entire group of catalysts (4 M, 8 M, and 10 M). Moreover, it can be observed that an increase in molarity leads to a number of defects that are progressively independent of the aging temperature. As a result,  $T_{10\%}$  shows a considerable variability

at 4 M, according to the aging temperature, while it is almost constant at 10M, where neither the defects nor the CO oxidation activity are influenced by the aging temperature.

A similar trend can be observed for the NO oxidation activity. In this case,  $T_{30\%}$  (that is, the temperature at which  $\text{NO}_2/\text{NO}_x$  is equal to 30%) is compared with the number of vacancies at 400 °C, which is a value that falls within the range of plotted representative temperatures for NO oxidation. The same conclusions that were previously reported for CO can be drawn, although the slope of  $T_{30\%}$  is lower than that of  $T_{10\%}$ , with reference to CO oxidation (at an equal  $\Delta T$  scale in the two sets of graphs).

Finally, the third set of plots refers to soot oxidation: both  $\text{O}_2$ - and  $\text{NO}_x$ -mediated reactions, expressed as  $T_{10\%}$  in Table 4, are reported, and are compared with the abundance of defects at 500 °C. The relationship between activity and defects is much weaker than in the other two cases (CO and NO, which are gaseous reactants), presumably because of the poor contact between the catalyst and soot in loose conditions. In this regard, it is known that the morphology of a catalyst plays a relevant role in the degree of contact between the catalyst itself and soot, although this effect is not easy to predict [19] or optimize [37,38]. However, it is interesting to note that when NO is introduced, the slope of the  $T_{10\%}$  curve generally increases, although marginally, as the NO oxidation promotes a reduction in  $T_{10\%}$  which, as previously mentioned, is influenced by the abundance of defects.

### 3. Materials and Methods

#### 3.1. Catalyst Preparation

A set of nine  $\text{CeO}_2$  samples was prepared, via a hydrothermal procedure, using different temperatures (150, 170, and 180 °C) and NaOH concentrations (4, 8, and 10 M) [18,37]. During a typical synthesis, appropriate amounts of NaOH and  $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$  (99%, Sigma-Aldrich, Steinheim, Germany) were dissolved in 70 mL and 10 mL of bi-distilled water, respectively. The two solutions were then mixed together and stirred for 1 h, and a milky slurry was thus obtained.

The final mixture was transferred to an autoclave (volume = 150 mL), which was 75% filled with deionized water. The mixture was aged at 150, 170, or 180 °C for 24 h. The precipitate was then centrifuged, washed and dried at 70 °C overnight. Finally, the powder was calcined at 550 °C for 4 h.

#### 3.2. Catalyst Characterization

The powder X-ray diffraction patterns were collected on an X'Pert Philips PW3040 diffractometer using  $\text{Cu K}\alpha$  radiation ( $2\theta$  range =  $20^\circ$ – $70^\circ$ ; step =  $0.05^\circ$   $2\theta$ ; time per step = 0.2 s). The diffraction peaks were indexed according to the Powder Data File database (PDF-2 1999, International Centre of Diffraction Data, PA, USA).

The Specific Surface Area ( $S_{\text{BET}}$ ) and total pore volume ( $V_p$ ) were measured by means of  $\text{N}_2$  physisorption at  $-196^\circ\text{C}$  (Micromeritics Tristar II 3020, v1.03, Micromeritics Instrument Corp., Norcross, GA, USA, 2009) on samples previously outgassed at 200 °C for 4 h. The specific surface area of the samples was calculated using the Brunauer-Emmett-Teller (BET) method.

The morphology of the samples was investigated by means of field emission scanning electron microscopy (FESEM Zeiss MERLIN, Gemini-II column, Oberkochen, Germany).

The reducibility of the catalysts was analyzed by means of CO-TPR (Temperature-Programmed Reduction). Amounts of 45 mg of powdered catalyst and 150 mg of silica were introduced into a quartz U-tube reactor. The reduction was started by flowing  $200\text{ mL min}^{-1}$  gas containing 2000 ppm of CO in  $\text{N}_2$  through the reactor. The temperature increased from room temperature to 700 °C at a programmed rate of  $10^\circ\text{C min}^{-1}$ .

Temperature dependent Raman measurements were performed using a Linkam TS1500 stage, under a 5x objective of a Renishaw InVia Reflex micro-Raman spectrometer, with laser excitation at 514.5 nm in backscattering light collection mode. The samples were kept at each temperature for 15 min before the spectra were acquired in order to reach thermal equilibrium. The  $\text{O}_v/\text{F}_{2g}$  value, which is

representative of the density of the oxygen vacancies, was calculated as the ratio of the integrated areas of the  $O_v$  and  $F_{2g}$  Raman peaks obtained by means of the deconvolution procedure.

### 3.3. Catalytic Activity Test

Catalytic tests were executed in a quartz U-tube reactor, with an inner diameter = 4 mm, heated by an electric furnace. The temperature was measured by means of a thermocouple placed in the middle of the catalytic bed. Each test was repeated to ensure the reproducibility of the obtained results. Temperatures corresponding to 10%, 50%, and 90% conversion (denoted as  $T_{10\%}$ ,  $T_{50\%}$ , and  $T_{90\%}$ , respectively) were taken as indices of the activity of the tested catalysts.

CO oxidation reaction: 100 mg of powdered catalyst was introduced into the reactor. The catalytic bed was pre-treated in He (flow rate = 50 mL  $\text{min}^{-1}$ ) for 1 h at 150 °C. The test was started by introducing 50 mL  $\text{min}^{-1}$  gas, containing 1000 ppm of CO and 10%  $O_2$  in  $N_2$ , into the reactor (gas hourly space velocity of 19,100  $\text{h}^{-1}$ ). The temperature was then raised by 5 °C  $\text{min}^{-1}$  over the 50 to 600 °C range.

Soot oxidation reaction: The catalytic bed was prepared by mixing 5 mg of carbon soot (Printex-U), 45 mg of powdered catalyst and 150 mg of inert silica in “loose” contact conditions. The catalyst-soot mixture was prepared by gently stirring it with a spatula for 60 s. This procedure appears sufficient to homogenize the mixture, even though the two solids are placed loosely in contact. The test was started by flowing 100 mL  $\text{min}^{-1}$  gas, comprising 50% of air and 50% of  $N_2$ , through the reactor. The temperature was increased from 100 to 700 °C at a 5 °C  $\text{min}^{-1}$  heating rate.

NO oxidation reaction: the catalytic bed was made up of 150 mg of inert silica and 45 mg of catalyst, gently mixed with a spatula. The catalyst was pre-treated with 100 mL  $\text{min}^{-1}$  of air at 100 °C for 30 min. The reactor was then cooled to 50 °C, and this was followed by the introduction of the NO flow (100 mL  $\text{min}^{-1}$  of a mixture of 550 ppmv NO and 10%-v  $O_2$  in  $N_2$ ) till the gas concentrations were stable. The reactor was heated from 50 to 650 °C at a programmed rate of 5 °C  $\text{min}^{-1}$ .

$NO_x$ -assisted soot oxidation reaction: the catalytic bed contained silica (150 mg), a catalyst (45 mg) and 5 mg of soot (Printex-U from Degussa), gently mixed with a spatula for 3 min. The catalyst was pre-treated with 100 mL  $\text{min}^{-1}$  of 20%-v  $O_2$  in  $N_2$  at 100 °C for 30 min. The reactor was then ramp-cooled (5 °C  $\text{min}^{-1}$ ) to 50 °C, and this was followed by the introduction of the NO flow (100 mL  $\text{min}^{-1}$  of a mixture of 550 ppmv NO and 10%-v  $O_2$  in  $N_2$ ). After stabilization of the  $NO_x$  concentrations, the reactor was heated gradually at a rate of 5 °C  $\text{min}^{-1}$  to 700 °C.

## 4. Conclusions

In the present work, a set of nanostructured ceria catalysts was prepared, by means of a hydrothermal procedure, using different synthesis conditions (i.e., temperature, pH values). As a result, samples with different shapes and structural properties were obtained and their catalytic activity was tested considering various oxidation reactions (namely, CO oxidation, NO oxidation,  $NO_x$ -free soot oxidation,  $NO_x$ -assisted soot oxidation). The 150 °C-4 M catalyst exhibited the best performances for both CO and NO oxidation, thus confirming the beneficial role of the textural properties on the overall activity on this set of  $CeO_2$  nanostructures. In fact, a higher surface area means more abundant surface defects that are in turn spatially accessible to reactant molecules. On the other hand, the 150 °C-8 M sample appeared to be the most promising catalyst for the soot oxidation reaction carried out in both the absence and presence of NO.

The role of the structural properties on catalytic activity was also investigated. High temperature in-situ micro-Raman measurements enabled the study of surface defects under the conditions involved in the catalytic processes. A direct (positive) correlation was observed between CO oxidation activity and the structural defects of the catalysts ( $O_v/F_{2g}$  ratios evaluated at 300 °C): the more defective the catalyst surface was, the greater the oxidation activity.

Similarly, a positive trend appeared for the NO oxidation activity ( $O_v/F_{2g}$  ratios evaluated at 400 °C). Interestingly, these correlations originated from the dependence of the defect density on the

specific surface area of the catalyst, instead of being related to the type of exposed crystal planes. On the other hand, a weaker correlation was achieved between the soot oxidation activity and the presence of surface defects, likely because of the loose soot-catalyst contact conditions. However, the soot oxidation activity appeared to be slightly more influenced by the abundance of surface defects in the presence of NO (NO<sub>x</sub>-assisted soot oxidation).

**Author Contributions:** M.P. and S.B. conceived and designed the experiments; M.D. performed the sample synthesis, XRD, BET, and catalytic activity tests; T.A. analyzed the catalytic data, while C.N. performed and analyzed the temperature dependent Raman spectroscopy measurements; N.R., D.F., and F.G. supervised the work.

**Conflicts of Interest:** The authors declare no conflict of interest.

## References

1. Trovarelli, A.; Fornasiero, P. *Catalysis by Ceria and Related Materials*, 2nd ed.; Imperial College Press: London, UK, 2013.
2. Heck, R.M.; Farrauto, R.J.; Gulati, S.T. *Catalytic Air Pollution Control: Commercial Technology*, 3rd ed.; Wiley-VCH: Hoboken, NJ, USA, 2006.
3. Ertl, G.; Knözinger, H.; Schüth, F.; Weitkamp, J. *Handbook of Heterogeneous Catalysis*, 2nd ed.; Wiley-VCH: Weinheim, Germany, 2008.
4. Bueno-López, A. Diesel soot combustion ceria catalysts. *Appl. Catal. B Environ.* **2014**, *146*, 1–11. [[CrossRef](#)]
5. Burch, R. Knowledge and Know-How in Emission Control for Mobile Applications. *Catal. Rev.* **2004**, *46*, 271–334. [[CrossRef](#)]
6. Russell, A.; Epling, W.S. Diesel Oxidation Catalysts. *Catal. Rev.* **2011**, *53*, 337–423. [[CrossRef](#)]
7. Eastwood, P. *Critical Topics in Exhaust Gas Aftertreatment*; Research Studies Press Ltd.: Baldock, UK, 2000.
8. Duprez, D.; Cavani, F. *Handbook of Advanced Methods and Processes in Oxidation Catalysis*; Imperial College Press: London, UK, 2014.
9. Fino, D.; Bensaid, S.; Piumetti, M.; Russo, N. A review on the catalytic combustion of soot in Diesel particulate filters for automotive applications: From powder catalysts to structured reactors. *Appl. Catal. A Gen.* **2016**, *509*, 75–96. [[CrossRef](#)]
10. Piumetti, M.; Bensaid, S.; Russo, N.; Fino, D. Nanostructured ceria-based catalysts for soot combustion: Investigations on the surface sensitivity. *Appl. Catal. B Environ.* **2015**, *165*, 742–751. [[CrossRef](#)]
11. Iijima, S. Helical microtubules of graphitic carbon. *Nature* **1991**, *354*, 56–58. [[CrossRef](#)]
12. Tans, S.J.; Verschueren, A.R.M.; Dekker, C. Room-temperature transistor based on a single carbon nanotube. *Nature* **1998**, *393*, 49–52.
13. Fan, S.; Chapline, M.G.; Franklin, N.R.; Tomblor, T.W.; Cassell, A.M.; Dai, H. Self-Oriented Regular Arrays of Carbon Nanotubes and Their Field Emission Properties. *Science* **1999**, *283*, 512–514. [[CrossRef](#)] [[PubMed](#)]
14. Geim, A.K.; Novoselov, K.S. The rise of graphene. *Nat. Mater.* **2007**, *6*, 183–191. [[CrossRef](#)] [[PubMed](#)]
15. Fedlheim, D.L.; Foss, C.A. *Metal Nanoparticles: Synthesis, Characterization, and Applications*; Marcel Dekker: New York, NY, USA, 2002.
16. Thomas, J.M.; Raja, R. Mono-, Bi- and Multifunctional Single-Sites: Exploring the Interface between Heterogeneous and Homogeneous Catalysis. *Top. Catal.* **2010**, *53*, 848–858. [[CrossRef](#)]
17. Agarwal, S.; Zhu, X.; Hensen, E.J.M.; Lefferts, L.; Mojet, B.L. Defect Chemistry of Ceria Nanorods. *J. Phys. Chem. C* **2014**, *118*, 4131–4142. [[CrossRef](#)]
18. Agarwal, S.; Lefferts, L.; Mojet, B.L. Ceria Nanocatalysts: Shape Dependent Reactivity and Formation of OH. *ChemCatChem* **2013**, *5*, 479–489. [[CrossRef](#)]
19. Piumetti, M.; van der Linden, B.; Makkee, M.; Miceli, P.; Fino, D.; Russo, N.; Bensaid, S. Contact dynamics for a solid–solid reaction mediated by gas-phase oxygen: Study on the soot oxidation over ceria-based catalysts. *Appl. Catal. B Environ.* **2016**, *199*, 96–107. [[CrossRef](#)]
20. Baek, S.O.; Field, R.A.; Goldstone, M.E.; Kirk, P.W.; Lester, J.N.; Perry, R. A review of atmospheric polycyclic aromatic hydrocarbons: Sources, fate and behavior. *Water Air Soil Pollut.* **1991**, *60*, 279–300. [[CrossRef](#)]
21. Laha, S.C.; Ryoo, R. Synthesis of thermally stable mesoporous cerium oxide with nanocrystalline frameworks using mesoporous silica templates. *Chem. Commun.* **2003**, *125*, 2138–2139. [[CrossRef](#)]

22. Taniguchi, T.; Watanabe, T.; Sugiyama, N.; Subramani, A.K.; Wagata, H.; Matsushita, N.; Yoshimura, M. Identifying defects in ceria-based nanocrystals by UV resonance Raman spectroscopy. *J. Phys. Chem. C* **2009**, *113*, 19789–19793. [[CrossRef](#)]
23. Wu, Z.; Li, M.; Howe, J.; Meyer, H.M., III; Overbury, S.H. Probing defect sites on CeO<sub>2</sub> nanocrystals with well-defined surface planes by Raman spectroscopy and O<sub>2</sub> adsorption. *Langmuir* **2010**, *26*, 16595–16606. [[CrossRef](#)] [[PubMed](#)]
24. Spanier, J.E.; Robinson, R.D.; Zhang, F.; Chan, S.W.; Herman, I.P. Size-dependent properties of CeO<sub>2-y</sub> nanoparticles as studied by Raman scattering. *Phys. Rev. B* **2001**, *64*, 245407. [[CrossRef](#)]
25. Weber, W.H.; Hass, K.C.; McBride, J.R. Raman study of CeO<sub>2</sub>: Second-order scattering, lattice dynamics, and particle-size effects. *Phys. Rev. B* **1993**, *48*, 178–185. [[CrossRef](#)]
26. Filtschew, A.; Hofmann, K.; Hess, C. Ceria and Its Defect Structure: New Insights from a Combined Spectroscopic Approach. *J. Phys. Chem. C* **2016**, *120*, 6694–6703. [[CrossRef](#)]
27. Su, Z.; Sha, J.; Pan, G.; Liu, J.; Yang, D.; Dickinson, C.; Zhou, W. Temperature-dependent raman scattering of silicon nanowires. *J. Phys. Chem. B* **2006**, *110*, 1229–1234. [[CrossRef](#)] [[PubMed](#)]
28. Zhao, Z.; Elwood, J.; Carpenter, M.A. Phonon Anharmonicity of PdO Studied by Raman Spectrometry. *J. Phys. Chem. C* **2015**, *119*, 23094–23102. [[CrossRef](#)]
29. Popović, Z.V.; Dohčević-Mitrović, Z.; Šćepanović, M.; Grujić-Brojčin, M.; Askrabić, S. Raman scattering on nanomaterials and nanostructures. *Ann. Phys.* **2011**, *523*, 62–74. [[CrossRef](#)]
30. Pandey, S.D.; Singh, J.; Samanta, K.; Sharma, N.D.; Bandyopadhyay, A.K. Temperature dependent variations of phonon interactions in nanocrystalline cerium oxide. *J. Nanomater.* **2015**, *16*, 154.
31. Popović, Z.V.; Dohčević-Mitrović, Z.; Cros, A.; Cantarero, A. Raman scattering study of the anharmonic effects in CeO<sub>2-y</sub> nanocrystals. *J. Phys. Condens. Matter* **2007**, *19*, 496209. [[CrossRef](#)]
32. Li, Y.; Wei, Z.; Gao, F.; Kovarik, L.; Baylon, R.A.L.; Peden, C.H.F.; Wang, Y. Effect of Oxygen Defects on the Catalytic Performance of VO<sub>x</sub>/CeO<sub>2</sub> Catalysts for Oxidative Dehydrogenation of Methanol. *ACS Catal.* **2015**, *5*, 3006–3012. [[CrossRef](#)]
33. Bensaid, S.; Piumetti, M.; Novara, C.; Giorgis, F.; Chiodoni, A.; Russo, N.; Fino, D. Catalytic Oxidation of CO and Soot over Ce-Zr-Pr Mixed Oxides Synthesized in a Multi-Inlet Vortex Reactor: Effect of Structural Defects on the Catalytic Activity. *Nanoscale Res. Lett.* **2016**, *11*, 494. [[CrossRef](#)] [[PubMed](#)]
34. Mamontov, E.; Egami, T.; Pennsylv, V.; Brezny, R.; Koranne, M.; Grace, W.R.; Grace, C.; Da, V.; Tyagi, S. Lattice Defects and Oxygen Storage Capacity of Nanocrystalline Ceria and Ceria-Zirconia. *J. Phys. Chem. B* **2000**, *104*, 11110–11116. [[CrossRef](#)]
35. Royer, S.; Duprez, D. Catalytic Oxidation of Carbon Monoxide over Transition Metal Oxides. *ChemCatChem* **2011**, *3*, 24–65. [[CrossRef](#)]
36. Mai, H.-X.; Sun, L.-D.; Zhang, Y.-W.; Si, R.; Feng, W.; Zhang, H.-P.; Liu, H.-C.L.; Yan, C.-H. Shape-Selective Synthesis and Oxygen Storage Behavior of Ceria Nanopolyhedra, Nanorods, and Nanocubes. *J. Phys. Chem. B* **2005**, *109*, 24380–24385. [[CrossRef](#)] [[PubMed](#)]
37. Miceli, P.; Bensaid, S.; Russo, N.; Fino, D. CeO<sub>2</sub>-based catalysts with engineered morphologies for soot oxidation to enhance soot-catalyst contact. *Nanoscale Res. Lett.* **2014**, *9*, 254. [[CrossRef](#)] [[PubMed](#)]
38. Bensaid, S.; Russo, N.; Fino, D. CeO<sub>2</sub> catalysts with fibrous morphology for soot oxidation: The importance of the soot-catalyst contact conditions. *Catal. Today* **2013**, *216*, 57–63. [[CrossRef](#)]

