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Review

A Comprehensive Review on Raman Spectroscopy Applications

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Abstract: Raman spectroscopy is a very powerful tool for material analysis, allowing for exploring the properties of a wide range of different materials. Since its discovery, Raman spectroscopy has been used to investigate several features of materials such carbonaceous and inorganic properties, providing useful information on their phases, functions, and defects. Furthermore, techniques such as surface and tip enhanced Raman spectroscopy have extended the field of application of Raman analysis to biological and analytical fields. Additionally, the robustness and versatility of Raman instrumentations represent a promising solution for performing on-field analysis for a wide range of materials. Recognizing the many hot applications of Raman spectroscopy, we herein overview the main and more recent applications for the investigation of a wide range of materials, such as carbonaceous and biological materials. We also provide a brief but exhaustive theoretical background of Raman spectroscopy, also providing deep insight into the analytical achievements.

Keywords: Raman spectroscopy; SERS; qualitative analysis; material characterization



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1. Introduction

Non-destructive spectroscopic techniques represent the top-choice for any kind of process monitoring [1,2].

Among all of the available techniques, Raman spectroscopy is one of the most solid and versatile tools to analyze several materials, both in lab and on-field conditions [3]. Raman spectroscopy was first independently developed in the first half of the 20th century by the Nobel laureate Chandrasekhara Venkata Raman [4] and Grigorij Samuilovič Landsberg [5], but it was established after the implementation of laser light equipped spectrometers in the second half of the century [6,7].

The establishment of Raman spectroscopy opened the path to a more detailed knowledge about materials, with a particular emphasis on carbonaceous materials such as graphite [8].

Throughout the years, Raman analysis has grown, reaching several industrial sectors such the food [9] and textiles sectors [10].

Raman spectroscopy displays several advantageous features over other techniques like infrared spectroscopy. For example, the quality of the signal collected is barely affected by the presence of water, allowing for use in plenty of applications where infrared analyses are not reliable [11]. A representative case study is the in-situ monitoring of a fermentative process where Raman techniques outperformed any other spectroscopic approach [12]. Nonetheless, Raman analysis suffers from some difficulties such as the

challenge of developing quantitative robust and trustworthy methods of data analysis [13]. Furthermore, the presence of highly active Raman species such as carbon particles could mask the presence of other species [14]. Several studies have been devoted to overcoming these drawbacks [15,16].

In this review, we overview the applications of Raman spectroscopy for the characterization of a wide range of materials. We also discuss in detail the phenomena behind the Raman scattering, thus providing a theoretical background for readers. Due to the immensity of the possible applications of Raman spectroscopy, we believe that a comprehensive overview of the state of the art of this technique is an essential asset to spread its knowledge.

2. Raman Spectroscopy: Principles and Instrumentation

2.1. Theoretical Principles: A Quantum Mechanical Approach

Electromagnetic radiation interacts with matter through absorption, transmittance, and/or scattering phenomena. An absorption process requires the matching of the incident photon's energy with the energy gap between the two electronic energy levels [17]. Conversely, the process of scattering does not require the presence of adequate energy levels, as this mechanism occurs when a photon interacts with a crystal lattice or a molecule, inducing a distortion of its electron cloud and changing the species polarization involving virtual states [18]. Nevertheless, the virtual state is a short lived one and it decays leaving the electron in the real electronic level of the system while the photon departs from the system. If the energy of the scattered photon matches the incoming one and the electron involved returns in a state with the same energy as the initial one, the scattering is elastic (termed Rayleigh—see Figure 1), otherwise it is inelastic. During inelastic scattering, the loss or gain of the photon energy is equal to the energy difference between the initial and final electronic levels. If the outgoing photon has a lower energy than the incoming one, the scattering is a Stokes one (Figure 1); in the opposite case, the scattering is an anti-Stokes one (Figure 1). The energy difference between the incoming photon and the outgoing one is called "Raman shift".

From a quantum mechanical approach, the energy corresponding to each vibrational level for a diatomic molecule is described by the following:

$$E_v = hv \left(n + \frac{1}{2} \right) \quad (1)$$

where v is the frequency of the vibration, h is Planck's constant, and n is the vibrational quantum number with integer values. For the sake of simplicity, the harmonic approximation expression is presented, but it has its limits. Chiefly, it only allows for fundamental transitions ($\Delta v = \pm 1$) and the separation between adjacent levels is supposedly constant, which is not the case for a real diatomic molecule [19]. From Equation (1), it can be seen that for $n = 0$, the energy is $E = \frac{hv}{2}$, a direct consequence of the Heisenberg principle.

The matrix element that defines the transition probability for a single Raman event between two electronic states is the same for the Stokes and anti-Stokes processes. The difference in the number of scattered photons is due to the different electron populations at the thermal equilibrium of the two electronic levels involved in the Raman process.

At room temperature, the majority of molecules populate the ground state, and the anti-Stokes components are much less intense than the Stokes ones. The different populations of the energy levels are described by the Boltzmann equation

$$\frac{N_{ex}}{N_g} = \frac{g_{ex}}{g_g} e^{-\frac{(E_{ex}-E_g)}{kT}} \quad (2)$$

where N_{ex} and N_g are the number of molecules in the excited state and in the ground state of energies E_{ex} and E_g , respectively, and g is the degeneracy of the energy level.

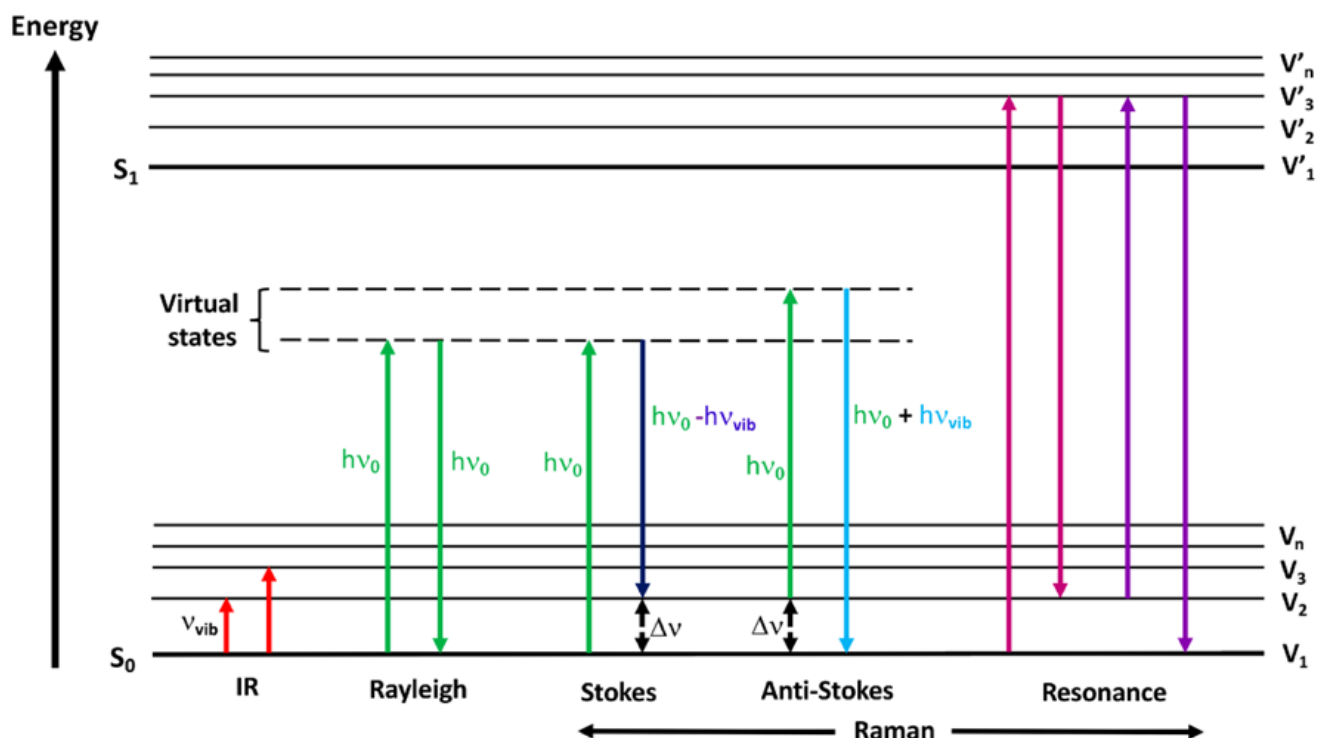


Figure 1. Jablonski energy diagram showing the transitions involved during infrared absorption, Rayleigh, Raman Stokes, anti-Stokes, and resonance Raman scattering as reported by Geraldès [20]. Reprinted under CC BY 4.0.

If the energy of the incoming photon is such that the electron can reach a real electronic state of the molecule, the Raman is called resonant (see Figure 1), and enhancements of up to 10^6 in signal intensity have been observed [21].

It should be emphasized that Raman scattering is a particularly weak phenomenon compared with Rayleigh, and only one in 10^8 scattered photons will undergo Raman scattering. On that account, all Raman spectrometers are equipped with some kind of device (usually a notch filter) to remove the Rayleigh component.

The intensity of the Raman scattered radiation is described by the following [22]:

$$I_R \propto I_0 v^4 N \left(\frac{\partial \alpha}{\partial Q} \right)^2 \quad (3)$$

where I_0 and v are the intensity and the frequency of the incident light, respectively; N is the number of scattering molecules in a given state; α is the polarizability; and Q is the amplitude of the vibrational coordinate. The last term on the right side shows that Raman scattering can occur only if a change in polarizability occurs due to the interaction with the incoming photon.

A misinterpretation of Equation (3) could lead to the fallacious conclusion that the best signal-to-noise ratio in Raman spectroscopy could be obtained by simply maximizing the frequency of the incident light. Unfortunately, as the beam of light on the sample becomes more energetic, the photon absorption and/or the material photodegradation increases substantially. For this reason, visible light lasers are commonly used. However, fluorescent phenomena could be very intense, significantly rising the background noise.

It is possible to enhance the Raman signal by exploiting the plasmonic effects induced by the presence of metal nanoparticles. This technique, known as surface-enhanced Raman spectroscopy (SERS), represents a powerful, highly specific, and sensitive technique for the investigation of the molecular structure at the single-molecular level based on Raman scattering. In the next section, the SERS theoretical background will be discussed.

2.2. Theoretical Background of Surface Enhanced Raman Spectroscopy Techniques

SERS is mainly exploited through the interaction between the surface of plasmonic nanostructures and the molecules adsorbed on it. Often, noble metals or copper nanoparticles [23] are used due to their resonant interaction with the electromagnetic waves, inducing localized surface plasmons covering almost all of the visible and near infrared wavelength range [24], as summarized in Figure 2.

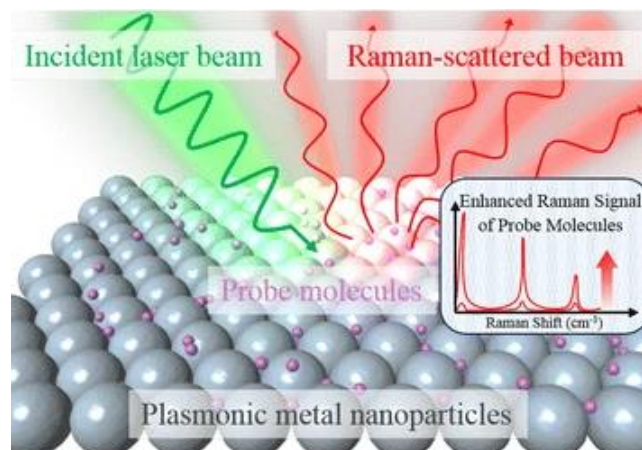


Figure 2. Schematic diagram of the SERS process. Reprinted from Jeon et al. [25]. Reprinted under CC BY 4.0.

The optical properties of the noble metal nanoparticles are dominated by strong features in the extinction spectra. The localized surface plasmons can be so intense that they significantly modify the local electromagnetic field [26]. This phenomenon could be explained using two different theories—the electromagnetic and chemical theories. The electromagnetic theory has strong experimental support [27] and will be the only one discussed in this section.

In order to explain the SERS dynamics, the optical response of a metal nanoparticles has to be analyzed according to Mie's theory [28,29]. A crucial parameter in Mie's theory is the complex dielectric function of the nanoparticle, which is calculated by the phenomenological Drude theory [30]. The dielectric function has its contributions from interband ($\epsilon_{inter}(\omega)$) and intraband ($\epsilon_{intra}(\omega)$) electron transitions, and can be written as follows:

$$\epsilon_{EXP}(\omega) = \epsilon_{inter}(\omega) + \epsilon_{intra}(\omega) \quad (4)$$

Interband contributions are due to electron transitions from occupied to empty states sitting in bands separated by an energy gap. Intraband contributions come from electron transitions at the Fermi level in incompletely filled bands. This contribution to the $\epsilon(\omega)$ is described by the Drude model, including an extra damping term τ . In the case of nanoparticles, the damping term is mainly related to the effect of the nanoparticle surface. The electron motion is altered by the NP surface, because the electron's mean free path λ , related to the lifetime τ , is comparable or larger than the particle size. The effect of the surface not only depends on the particle size, but also on its shape [31]. In order to include surface dispersion in the Drude model, the intraband contribution is modified by changing the damping term. Finally, the dielectric function, which depends on the nanoparticles structure and includes the contribution of the free electrons, interband transition, and surface damping, is obtained as follows [32]:

$$\epsilon(\omega, a) = \epsilon_{inter}(\omega) + \epsilon_{intra}^{NP}(\omega, a) = (\epsilon_{exp}(\omega) - \epsilon_{intra}(\omega)) + \left(\frac{\omega_p^2}{\omega \left(\omega + \frac{i}{\tau} + \frac{i}{\tau(a)} \right)} \right) \quad (5)$$

where ω_p is the plasma frequency and $1/\tau$ is the damping constant due to the dispersion of the electrons. The surface dispersion of a sphere of radius, a , is given by $1/\tau$ (σ) = v_f/a , where v_f is the Fermi velocity of the electron cloud. Accordingly, the smaller the particle, the more important the surface dispersion effect.

Electromagnetic enhancement occurs when the incident light is in resonance with the surface plasmon modes of a metallic thin film or nanoparticle. In this case, it is crucial to fine-tune the spectral positions of the surface plasmons to achieve the resonance condition. When the electromagnetic field is induced on the metal surface, the Raman modes of the particle under analysis adsorbed on the metal surface are strongly enhanced. The average field of the Raman scattered light due to the described mechanism, given by the following Equation [33]:

$$E_R \propto \alpha_R A(v_L) E_0 \quad (6)$$

where E_0 is the magnitude of the incident field, $A(v_L)$ is the field enhancement averaged over the surface of the metal particle, and α_R is the component that takes into account the Raman scattering. Thus, the amplitude of the SERS-scattered field (E_{SERS}) is given by the following equation

$$E_{SERS} \propto \alpha_{SERS} A(v_L) A(v_S) E_0 \quad (7)$$

where $A(v_S)$ is the enhancement factor of the second enhancement process and α_{SERS} describes a cross-section of the new Raman process of the adsorbed molecule, and the average SERS intensity will be proportional to the square modulus of E_{SERS} . So, the SERS Stokes power (P^{SERS}) is given by

$$P^{SERS}(v_S) = N \alpha_{SERS} |A(v_L)|^2 |A(v_S)|^2 I(v_L) \quad (8)$$

where N is the number of molecules involved in the SERS process and $I(v_L)$ the excitation laser intensity.

As reported by Álvarez-Puebla [34], the Raman signal enhancement for a given excitation wavelength ($ER_{(\lambda)}$) is

$$ER_{(\lambda)} = \frac{|E_{out(\lambda)}|^2 - |E_{out(\lambda-\lambda_S)}|^2}{E_0^4} \quad (9)$$

where $E_{out(\lambda)}$ and $E_{out(\lambda-\lambda_S)}$ are the electromagnetic fields generated by the incident excitation and the Stokes's shifted Raman, respectively. Using Equation (9) for the estimation of the enhancement of the Raman signal is quite difficult, so it generally evaluated experimentally using the following equation

$$ER_{(\lambda)} = \frac{I_{SERS}}{I_{Raman}} f \quad (10)$$

where I_{SERS} is the intensity of SERS signal, I_{Raman} is the intensity of Raman signal, and f is ratio between the number of molecules in the Raman and SERS experiment.

According to Equations (9) and (10), the average SERS electromagnetic enhancements are of the order of $\sim 10^4$ – 10^6 , and are sometimes as high as $\sim 10^8$ (depending on the definition and measurement procedure). Maximum values for electromagnetic enhancement for isolated single colloidal silver and gold spheroids are of the order of 10^6 – 10^7 [35–37]. As discussed, the theory predicts a stronger enhancement of electromagnetic fields for sharp features and large curvature regions of the metal nanoparticles. For example, it was shown that the electromagnetic SERS enhancement factor could be increased up to nearly 10^{11} when the sphere degenerates and becomes sharper [38]. In addition, closely spaced interacting particles can provide extra field enhancement [39]. For example, electromagnetic enhancement factors up to 10^{11} have been estimated for the mid-point between two silver or gold spherical particles separated by a gap of 1 nm [38]. Furthermore, the metal nanoparticles exhibit fractal properties [40], such as colloidal clusters formed by aggregation of colloidal

particles or metal island films [41]. In such colloidal cluster structures, the individual dipole oscillators of the isolated nanoparticles are coupled and cover wide frequency region from the visible to the near infrared. In this case, the excitation is not distributed uniformly over the entire cluster, but tends to be spatially localized in some hot spots, and the surface of a fractal colloidal cluster structure shows a very inhomogeneous field distribution [42]. The size of the “hot spots” can be as small as a few nm. Their locations depend strongly on the geometry of the fractal object and on the excitation wavelength and polarization of the optical fields [43]. SERS enhancement is estimated to be up to 10^6 for spatially isolated small colloids and 10^7 – 10^8 for colloidal clusters [44].

The applications of SERS include sensing and imaging applications such as spectro-electrochemistry, single molecule SERS, and analytical and biological applications [45–48]. The power of SERS is its great ability to identify chemical species and to obtain accurate structural information regarding polymer and materials science, biochemistry and biosensing, catalysis, and electrochemistry [49,50]. Interestingly, SERS can also be used to investigate the orientation of single molecules, as reported by Moskovits et al. [51]. The authors observed the rise of normally forbidden vibrations in SERS spectra as a result of the fast spatial rate of the radiative electric field. Accordingly, they assumed that the change in Raman spectra was due to the different spatial arrangement of the adsorbed molecules on the SERS substrate. As exhaustively discussed by Aroca [52], this approach could be used for evaluating several parameters, ranging from structural to electrical.

Another interesting technique based on the enhancement of the Raman signal is tip enhanced Raman spectroscopy (TERS) [53].

TERS is a near-field method to spectroscopically analyze a wide range of chemical and biological samples with a high spatial resolution of a few nanometers, only limited by the tip apex size and shape (generally of diameter ~ 10 nm). The tip acts as an individual plasmonic nanostructure and is scanned over the sample using a scanning probe microscopy technique, as comprehensively explained by Fiederling et al. [54]. It provides a localized region of SERS enhancement when brought within a few nanometers from the sample [53]. In this way, it is possible to overcome the limitations of SERS enhancement due to its critical dependence on the substrate preparation. TERS offers the unique ability to attain enhanced Raman signals and topographical information simultaneously. Due to the highly localized nature of the probe, TERS provides signal-to-noise ratios of 10^4 – 10^5 and achieves an enhancement factor up to 10^6 [55]. It has been used to obtain high-quality Raman spectra of DNA pyrimidine bases [56], and for the nucleobases in a single RNA strand, it provided a high sensitivity at a lateral resolution, down to a few tens of nucleobases [57]. Investigations of cytochrome C revealed that the TERS spectra from different areas of biomacromolecules could be significantly different from those obtained with bulk Raman measurements of the sample [58].

The metallic tip can be replaced by a monolayer of metal nanoparticles, each coated with an ultrathin shell of silica or alumina. Each nanoparticle acts as a tip in the TERS system, and the use of a chemically inert shell coating around the nanoparticle protects the SERS-active nanostructure from contact with the sample, and can overcome the SERS limitation of the roughened metal substrates. SHINERS (shell-isolated nanoparticle-enhanced Raman spectroscopy) can obtain an enhancement that is two to three orders of magnitude higher than that for a single TERS tip, and its main advantage is a much higher detection sensitivity and several practical applications to a great amount of materials with diverse morphologies [59].

2.3. Instrumentation: Raman and microRaman

The first Raman spectroscopy approach employed a non-coherent light sources [4], but today all Raman spectrometers employ a laser source, either a continuous wave or a pulsed laser. Since few years ago, CW lasers were Kr, Ar, or Kr–Ar lasers, which supplied the sample with a continuous source of photons in the visible or IR range, but these have been largely replaced by diode-pumped solid-state laser [60]. Pulsed lasers are usually Nd:YAG (1064 nm) or excimer lasers, and are characterized by a much higher power output (10–100 mW) [19]. It should be noted that given the limited duration of the pulse, the

detector needs to be locked-in. Moreover, the use of ultrafast lasers combined with gated detectors allows for exploiting the difference in emission time between Raman scattering and fluorescence, substantially removing the latter from the recorded spectrum [61,62]. Different systems can be employed in Raman spectrometers in order to collect the spectral intensity of the signal. The dispersive method requires the use of a diffraction grating, composed of a series of grooves on a reflective support. By changing the angle of incidence of the scattered polychromatic radiation, a frequency selection is possible, thus measuring the intensity versus the frequency or wavenumber. Conversely, the use of an interferometric spectral analyzers allows for acquiring the whole spectrum. Then, by means of a computer algorithm, Fourier transformation is performed, switching from the time domain to the frequency domain. This technique is referred to as FT–Raman spectroscopy.

Conventional Raman spectroscopy has a relatively low spatial resolution ($\sim 1 \text{ mm}^2$), hence, for the characterization of nanostructured inorganic and organic materials alike, micro-Raman spectroscopy is preferred. Micro-Raman spectroscopy combines a Raman spectrometer with an optical microscope. By using a high magnification objective, the laser beam can be focused on an area of approximately $1 \text{ }\mu\text{m}$, being limited only by the diffraction limit of light, reaching a scanning sample depth ranging from few hundreds of nm up to $1 \text{ }\mu\text{m}$ [63,64].

Micro-Raman spectroscopy has found relevant applications in the evaluation of internal stresses in various nanocomposite materials [65–67], mapping phase and composition of polymer blends [68,69], probing structure and electronic properties of graphene [70], and in carbon nanotubes [71]. Another important use for μRS is in the determination of the crystal orientation of nanometric structures, such as CdS or GaN nanowires [72,73], which is particularly critical in the fields of electronic and optoelectronic materials.

3. Applications of Raman Spectroscopy

3.1. Carbon Based Materials

3.1.1. Carbon Nanotubes

Raman spectroscopy can be a powerful tool for the characterization of carbonaceous materials [74]. Accordingly, carbon nanotubes (CNT) have been deeply studied with Raman spectroscopy, helping to unravel the unique features of single-(SWCNTs), double (DWCNTs), and multiwalled (MWCNTs) carbon nanotubes. The one-dimensional (1D) confinement of electronic and phonon states is responsible for the optical and spectroscopic properties observed in SWNTs [75]. Every CNT has a unique Raman active mode, called a radial breathing mode (RMB), at around $100\text{--}200 \text{ cm}^{-1}$ [76], as reported in Figure 3.

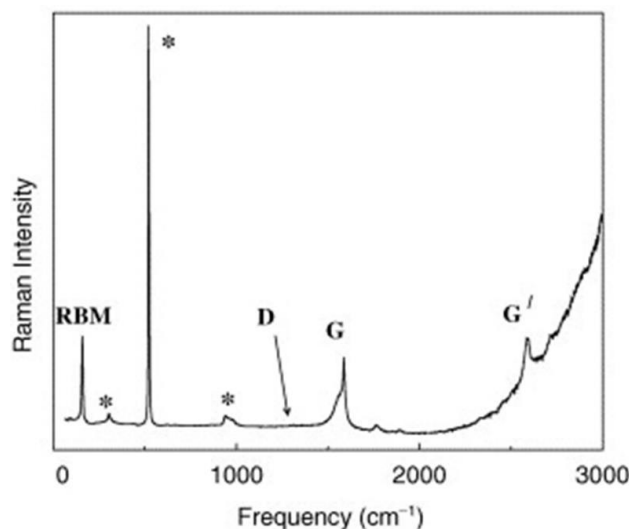


Figure 3. Raman spectra of isolated SWCNTs collected using a laser source with a wavelength of 785 nm [77]. Bands marked with * are due to the substrate of Si/SiO₂. Reprinted with permission from Elsevier.

RBM is crucial for the identification of the different types of CNTs geometry and represents proof the presence of SWNTs, while it is hardly detectable for other CNTs without using SERS [78]. RBM is inversely proportional on the CNT diameter through the equation Equation (11) [75], as follows:

$$\omega_{RBM} = \frac{A}{d + B} \quad (11)$$

where A and B are determined experimentally and d is the diameter of the CNT. If $d > 2 \mu\text{m}$, it is difficult to observe RBM because its intensity is very low. With d values obtained from ω_{RBM} and if the electronic transition energy (E_{ii})~the laser energy (E_L), it is possible to perform chiral indice (n,m) assignments for the isolated SWNCTs [79], the diameter d , the density of electronic states, the electronic energy, and the chiral angle θ [79]. Another important feature in the Raman spectra of SWNTs is the so-called G band spectra, which is split into two features around 1580 cm^{-1} (one peak at 1570 cm^{-1} and the other at 1590 cm^{-1}). The peak at 1570 cm^{-1} is also called G^- , and is associated with vibrations of the carbon atoms around the circumferential direction. The peak at 1590 cm^{-1} is called G^+ . G^+ corresponds to the longitudinal wave mode in the CNT axial direction, and G^- corresponds to the transverse wave mode perpendicular to the axis. The G^- Raman shift is proportional to the inverse square of the CNT diameter ($\propto 1/d^2$). The frequency of G^+ appears near 1590 cm^{-1} , regardless of the diameter of the carbon nanotube. The G band is derived from the optical vibration of two adjacent carbon atoms on the wall of the CNT [80], and its frequency does not depend on the excitation wavelength. It can be useful for distinguishing between semiconducting from metallic SWNTs and from the determination of d . Furthermore, the analysis of the G band could be a solid choice for evaluating the \dot{i} transfer of charge after SWCNT doping [75]. Concerning the characterization between metallic and semiconducting SWNTs, the following is know: for metallic nanotubes, the lower frequency component of the G peak (ω_G^-) has a wide and asymmetric profile, while for semiconducting CNTs, ω_G^- and ω_G^+ have a narrow profile that fits to a Lorentzian line shape [74]. Considering that G band is a first order process, the frequency ω_G^+ is independent of the diameter and chiral angle (θ) [75]. However, ω_G^- depends on d , but not on θ . There are other two features, called second order features [80], named the G' band (between 2775 and 2950 cm^{-1}) and the D band (1340 cm^{-1}), originating from a highly dispersive double resonance process [78]. The D peak is generally around 1350 cm^{-1} and derived from defective structures. The intensity ratio of the G band to D band (I_D/I_G ratio) is used as an indicator of the defect amounts in CNTs, increasing with their growth of defectiveness.

Their band frequencies strongly depend on d [79]. The presence of the G' band indicates the presence of a long-range order in the sample and is shown for MWCNTs with more than four layers [81]. G' and D are useful for studying the structural modification of the CNTs sidewalls. Kang al. [82] studied the D band dependence through heating and irradiation. They observed three components of the D band, called D^1 (1313 cm^{-1}), D^2 (1340 cm^{-1}), and D^3 (1355 cm^{-1}). If the laser irradiation increased, the D^1 and D^2 also intensity increased, while the D^3 intensity increased with heating in the air. These were due to the different defects introduced in the structure of the SWCNTs. The purity of the sample can also be investigated using the band intensity ratio I_D/I_G .

Concerning MWCNTs, the interpretation of their Raman spectra is generally more difficult than for SWCNTs [83]. Accordingly, its interpretation is often based on the results obtained for SWCNTs. In many cases, it is impossible to see the presence of RBM or the splitting of the G band, because the inner shell in MWCNTs has a diameter that is too much big. This proves that the MWCNT spectrum is similar to that of graphite, and shows little or no effects of cylindrical geometry. The G band of MWCNTs is made up of a broad range asymmetric features. With the use of polarized Raman scattering on an SERS substrate, Zhao et al. [84] observed the presence of RBM peaks at low frequency regions ($100\text{--}600 \text{ cm}^{-1}$). The inner diameter distribution calculated from the RBM frequencies coincided with the observation of high resolution transmission electron microscopy. The intensity of the modes decreases

if the dimension of the diameter increased, and when $d > 2 \mu\text{m}$, the modes disappeared. By using SERS, a splitting of the tangential stretching G band modes was found, proving that MWCNTs have different Raman spectra characteristics compared with other sp^2 carbon allotropes [83]. These results were reached by assigning a G band mode from the inner tube and a graphite-like mode from the outer cylinders. The wall–wall charge transfer is responsible for the relative frequencies of the modes. A splitting of the D band in two different features has also been found when the temperature is low [85]. Due to the carbon impurities effect, the ratio between I_D and I_G cannot be used to calculate the purity of MWCNT. The G' peak can be used in order to overcome this difficult, because the G' band intensity depends on the purity of the sample: a disorder structure would not allow for the coupling effect that is fundamental for the two-phonon process. This is due to the fact that the G' band results from a two phonons process and its intensity is highly sensitive to the sample purity. A ratio containing the G' peak at about 2700 cm^{-1} can represent a good measurement of the MWNT quality [86]. In the Raman spectra of MWCNTs bundles, two graphitic bands are visible—the G band at 1580 cm^{-1} [87] (associated to the in-plane vibration of the C-C bond) and the D band at around 1342 cm^{-1} . Rao et al. [88] showed that the D band of MWCNTs does not imply that the structure of the sample is disordered. Kumar et al. [89] analyzed the effect of the laser and energetic ion irradiation on the Raman spectroscopy. The ratio I_D/I_G decreases if the laser power density increases. The authors reported the same results with ion irradiation: the ratio between D and G peak (I_D/I_G) decreases at a low ion influence and increases at a high ion flux. Raman spectroscopy can be useful in order to evaluate the number of carbon nanotube walls. Indeed, Chaunchaiyakul et al. [90] demonstrated that the G' and G band intensity ratios are correlated with the number of walls of the carbon nanotubes. The G' band intensity can increase thanks to the inner-tube interaction between the carbon nanotubes. This phenomenon is due to the quantum interference between Raman scattering paths resulting from the interaction between the walls of the carbon nanotubes.

3.1.2. Graphene and Graphene-Like Materials

Over the years, graphene has found plenty interest thanks to its properties, like is mechanical, thermal, optical, and electrical properties. It is a two-dimensional material of sp^2 hybridized carbon atoms [91] and has a hexagonal lattice with conjugate covalent bonds formed between the two adjacent carbon atoms [92]. We can distinguish several types of graphene and graphene-like materials, named monolayer graphene, bilayer graphene (2-LG), tri-layer graphene, and multilayer graphene, with a structure resembling that of graphite [93]. The Raman spectra of pure single graphene displays two important peaks named G (at around 1582 cm^{-1}) and 2D (at around 2700 cm^{-1}) [92] (Figure 4).

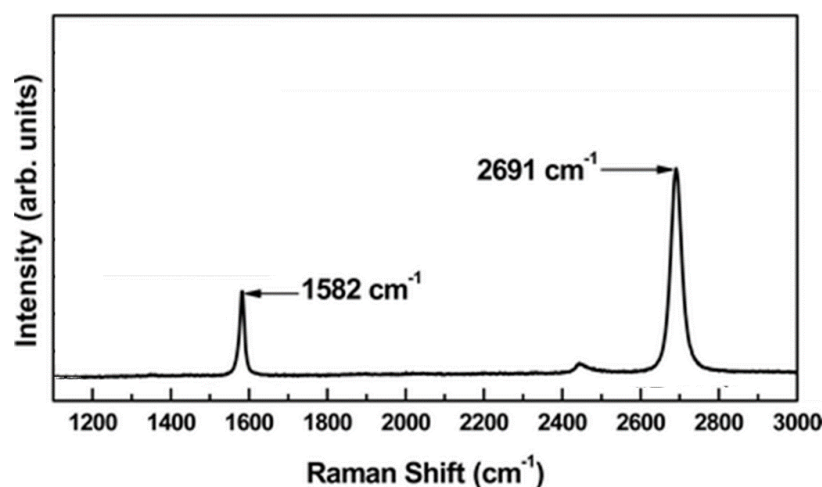


Figure 4. Raman spectra of single layer highly pure graphene as reported by Calizo et al. [94] Reprinted with all permissions from American Chemical Society (Copyright © 2007 American Chemical Society).

The G band can be used for the analysis of the structure of the graphene lattice that can be of two types: zigzag or armchair. Defective graphene displays a D peak arising from the presence of boards and defects. However, only armchair edges show the D peak because they are able to experience elastic scattering. We can see the D peak in the zigzag edges, but its intensity is very low, and its presence is due to the roughness and non-uniformity of the edge structure. The bands are also influenced by the temperature [95,96] with a split up of the G band and the G' band at a low temperature, and a downshift of the G band's frequency at a high temperature. These changes are reversible in thermal cycles with cooling and heating. Another parameter that can be used to calculate the number of graphene layers is the ratio between G' and G; this ratio $I_{G'}/I_G$ is low if the number of layers is high. In their work, Ferrari et al. [70] observed that the intensities of the D and G' peaks changed in position shape and intensity with the number of graphene layers. This is due to the evolution of the electronic structure and electron–phonon interactions. Regarding the G band, its intensity increases in a linear fashion if the number of layers is increased. Al-Hazmi et al. [97] used Raman spectroscopy in order to calculate the number of atomic planes of the graphene film grown on the Si substrate. This was possible thanks to the ratio between the intensity of the graphene G band and the optical phonon band of the Si substrate (I_G/I_{Si}). Koo et al. [98] used Raman spectroscopy in order to evaluate the role of poly(methyl methacrylate) (PMMA) on graphene fabricated by chemical vapor deposition. They compared bare PMMA-covered supported and PMMA-covered suspended graphene: their G and G' band positions were progressively downshifted. They also demonstrated, thanks to the mapping of G and G' shifts into strain and doping contributions, that PMMA residue exerts a tensile strain of about 0.15% on the graphene/substrate compared with that of bare graphene. Woo et al. [99] reported a method to measure the carrier mobility of chemical vapor deposited graphene from their Raman spectra. They found that the width of the G' peak decreased in proportion to the carrier mobility, while the ratio of I_D/I_G did not show changes. Furthermore, the mobility generated from the long-range scattering showed a linear relationship with the carrier mobility logarithm. For the monolayer graphene, one of the most important features in the Raman spectra is the G band at 1582 cm^{-1} . This is a normal first order Raman scattering process. There are also G' and D bands; the G' band frequency is twice the D band frequency and for this reason it is also called 2D—as they are second order process. Both the D and G' band frequencies change as a function of the laser intensity; the frequency of the D band changes in a linear way, while the upshift of the G' band is double that of the D band. The G' band has a single Lorentzian feature with a width of about 24 cm^{-1} ; this spectrum is much larger than the G band spectrum, which can reach four times the G band intensity [100]. Bilayer graphene shows four-peak G' components due to its electronic structure, and consists of two conduction bands and two valence bands. If we compare the peaks of mono-, bi-, and tri-layer graphene with highly oriented pyrolytic graphite, it is possible to observe that highly ordered graphite shows a higher intensity peak at higher frequencies. Regarding graphene growth on an SiC substrate, there is the problem for the Raman spectra: the SiC spectra (between 1450 and 1750 cm^{-1}) is superimposed on the G-band of the graphene. Moreover, Gokturk et al. [101] showed that during sequential exposure to hydrochloric acid vapors and ammonia of a poly(vinylchloride) supported graphene film, a shift can be appreciated in the Raman 2D band towards a lower and then higher wavenumber. This suggests n-type doping and restoration of graphene to its original state.

3.1.3. Amorphous and Disordered Carbon

Amorphous carbon has attracted much interest in previous years thanks to its wide-ranging properties, allowing for it to be used in many fields like plastic, textile, and health-care industries [102]. Depending on the chemical bonds formed, carbon atoms' orbitals can hybridize in three different ways— sp^3 , sp^2 , and sp . The properties of the amorphous carbon film are defined by the atoms and their arrangements. The sp^3 atoms are accounted for their mechanical properties, while the sp^2 atoms can determine the optical properties [103]. The atoms can coexist despite having different hybridizations. We can distinguish two

forms of amorphous carbon: the hydrogenated amorphous carbon, fabricated by plasma deposition of hydrocarbons, and the black semiconducting film, prepared via sputtering or evaporation [104]. Raman spectroscopy can excite the π and σ states, so it is able to detect sp^2 and sp^3 sites [105]. It is also used in order to evaluate the disorder of the carbon structures. The Raman spectrum depends on the sp^2/sp^3 ratio, bond disorder, presence of sp^2 rings or chains, and clustering of the sp^2 phase. The ordering of the sp^2 phase can be explained through the use of a three-stage model proposed by Ferrari et al. [105], with (i) a transition from graphite to nanocrystalline graphite, (ii) a transition from nanocrystalline graphite to amorphous carbon (a-C), and (iii) a final transition from a-C to tetrahedral amorphous carbon (ta-C). During these transitions, the G peak's position upshifts to higher values and the sp^3 content is enhanced. Amorphous carbon often displays a Raman spectrum with a large G band (1510 cm^{-1}) and D band (1350 cm^{-1}), which are sometimes overlapped. Furthermore, there is the so called T peak (at 1060 cm^{-1}), which is visible only with ultraviolet (UV) excitation [105]. The T peak is generated by the C-C sp^3 vibrations. Regarding the G peak, it originates from the bond stretching of sp^2 atoms in the rings and chains, while the D peak is due to the breathing modes of sp^2 atoms that are present in the ring [105]. If there are no rings, there is no D peak. G' is very low in intensity and is very broad, so it cannot be used in order to characterize amorphous carbon. The Raman spectrum of amorphous carbon is dominated by G and D bands, because the π states are of a lower energy than the σ sites, so they are much more polarizable and there is much scattering by the sp^2 sites [106]. The width of the G peak is useful in order to analyze the degree of disorder of the sample: the higher is the disorder, the higher the width. Ferrari et al. [105] defined G peak dispersion as the change of the G peak position for different excitation wavelengths. There is a linear variation for the G peak with different excitation wavelengths: this demonstrates that the higher is the disorder, the higher the dispersion. However, we can also observe a decrement of the Raman shift and shape changes at different wavelengths. The G peak dispersion is due to a selection of sp^2 configurations or clusters with a broad π band gap and, correspondingly, more elevated vibration frequencies. Furthermore, the dispersion can be used in order to discriminate two types of material: in the material with only sp^2 rings, there is a saturation of G peak dispersion at a maximum of 1600 cm^{-1} . However, if the materials contain sp^2 chains, the G peak can reach 1690 cm^{-1} and this is the case for ta-C or ta-C:H [105]. Concerning the clusters of sp^2 , they enhanced the G peak position in the Raman spectrum. On the other hand, if a UV light is used as a source of excitation, a decrease of the G peak position can be discerned if the number of clusters is high [105]. Thanks to this property, if two different samples have the same (or similar) Raman spectra, with UV excitation we can discriminate them as follows: the sample with the lower G peak position has more sp^2 clustering. The use of UV Raman is particularly useful for the investigation of hydrogenated amorphous carbon (a-C:H) [107], because for highly hydrogenated samples, the Raman spectra is overlapped by photoluminescence. UV Raman is able to overcome this problem and allows for measuring G and D peaks. The D band position is related to the type of amorphous carbon [107] and its intensity is due to the presence of six-fold aromatic rings. With the passage from amorphous carbon to tetrahedral amorphous carbon, the D peak disappears if a Breit Wigner Fano fit is used due to the increment of sp^3 sites, and so the ratio I_D/I_G is close to 0. During the transition from a-C to ta-C, there is also an increment of the amount of sp^3 sites (from 10–20% to 85%), while the sp^2 sites change from rings to chains [105]. There is also an increase in the G peak position (from 1510 cm^{-1} to 1570 cm^{-1}) with sp^3 content, and this is due to the change of sp^2 configuration from rings to olefinic groups. Tuinstra and Koenig [8] exposed an equation useful to evaluate the ratio of I_D/I_G in graphite structures, as reported in Equation (12):

$$\frac{C(\lambda)}{L_a} = \frac{I_D}{I_G} \quad (12)$$

where L_a is the graphite crystallite size and C is the Raman coefficient. This relation is not valid for amorphous carbon structures, because there are more defects and there is an increase of ordered rings. So, in this case the ratio become proportional to M, the number

of ordered rings. A new relationship was proposed by Ferrari et al. [105], as reported in Equation (13):

$$\frac{I_D}{I_G} = C'(\lambda)L_a^2 \quad (13)$$

Mariotto et al. [108] found that the ratio of I_D/I_G is enhanced if the annealing temperature is increased, and so there is a progressive increase in the graphitic domains. The I_D/I_G ratio in amorphous carbon is a measure useful to evaluate the dimension of the sp^2 phase that is organized in rings. The G peak dispersion is notable in more disordered carbon because in that case there are different configurations with different local band gaps and different phonon modes. In contrast, the D peak dispersion and the ratio of I_D/I_G decrease if there is a higher disorder. Wang et al. [109] found a correlation between the G peak and sp^3 atoms. They analyzed hydrogenated and H-free diamond-like carbon (DLC) samples. For the hydrogenated samples, the dispersion rate of the G peaks showed a linear correlation with sp^3 , while the FWHM of the G peak was non-linearly correlated with the sp^3 of the H-free DLC samples. Zhang et al. [110] proposed a method useful for calculating the intensity of the G band as a sum of three different Raman scattering intensities that derive from three different sp^2 clusters, as reported in Equation (14):

$$I_G(\omega, \lambda) = I_g(\omega, \lambda) + I_r(\omega, \lambda) + I_c(\omega, \lambda) \quad (14)$$

where ω is the Raman frequency; λ is the wavelength of the excitation laser; and I_g , I_r , and I_c are the intensities of the scattering of the nanocrystalline graphite, fused aromatic ring, and olefinic chain cluster, respectively. The ratio between the area of the T peak and the G peak I_T/I_G is useful in order to calculate the sp^3 content for ta-C [111]. Ta-C has a spectrum with two bands at around 1100 cm^{-1} and 1600 cm^{-1} , which are derived from sp^2 and sp^3 , respectively. A T peak of about 1060 cm^{-1} and a ratio of I_T/I_G of $\sim 0.40\text{--}0.42$ in H-free samples implies a sp^3 content of about 80% [105]. If the ratio of I_T/I_G is about $0.30\text{--}0.40$, the sp^3 content is about 60–80% [105]. Furthermore, a ratio of $I_T/I_G < 0.2$ indicates a poor sp^3 content (lower than 20–30%). Gilkes et al. [112] demonstrated that the 1100 cm^{-1} mode is derived from a mode of 1400 cm^{-1} that shifts to 1100 cm^{-1} when the sp^3 content is increased. If the structure contains atoms such H (a-C:H), a parameter called m/I_G can be used in order to calculate the amount of hydrogen bonded. In the ratio, m indicates the slope of the photoluminescence background in the range $800\text{--}1800\text{ cm}^{-1}$. However, Pardanaud et al. [113] demonstrated that the ratio of m/I_G is not very accurate because it can be influenced by various photoluminescence quenching processes and also because it is not sensitive to H bonded to C sp^2 . They used Raman spectroscopy to identify the H content in a-C. They found a band at 860 cm^{-1} attributed to H bonded to C (sp^2). The height ratio between D and G can be used in order to estimate the H content, because it is almost linear in the full range of H content. Casiraghi et al. [107] analyzed different a-C:H. They found two different behaviors. In the first case, the G peak width and G peak dispersion increased simultaneously, and the Raman parameters varied in the same way as in the hydrogen-free carbon film. In the second case (and this is the case of polymeric a-C:H), they found the opposite results, with the G peak dispersion varied in relation to the optical gap and hydrogen content, while the G peak width varied with the density. Liu et al. [114] studied the spectra of ta-C films as a function of substrate bias voltage. The higher the substrate bias, the higher the sp^3 atoms of the diamond-like carbon film (they tested films of about 70 nm), while for a ta-C of 2nm, they observed a linear decrease in sp^3 content when the bias was incremented. In another study by Liu et al. [115], they used the Raman spectral of ta-C as a function of the oblique angles of substrates. From the spectra, they observed that the higher the substrate tilting angle, the higher the sp^2 content, while less sp^3 content was observed. The substrate tilting depends on the sp^3 content and on the order of the sp^2 cluster, so from the Raman data, they observed a decrease of internal stress while there were no changes in hardness that depended on only the sp^3 content. Schwan et al. [116] found two additional peaks at 1180 cm^{-1} and 1490 cm^{-1} in the spectra

of the carbon film prepared by magnetron sputtering. They demonstrated that the peak at 1180 cm^{-1} was due to the presence of sp^3 stretching in the carbon film.

Another interesting material is the disordered carbon produced from the pyrolytic conversion of both natural and synthetic polymeric materials [117]. As shown in Figure 5, the Raman spectra of pyrolytic carbon displays a G and D peak with more components compared with graphite, but a profile that is not close enough to be similar to diamond-like materials. Accordingly, pyrolytic carbon could be described as a very highly defective graphitic material.

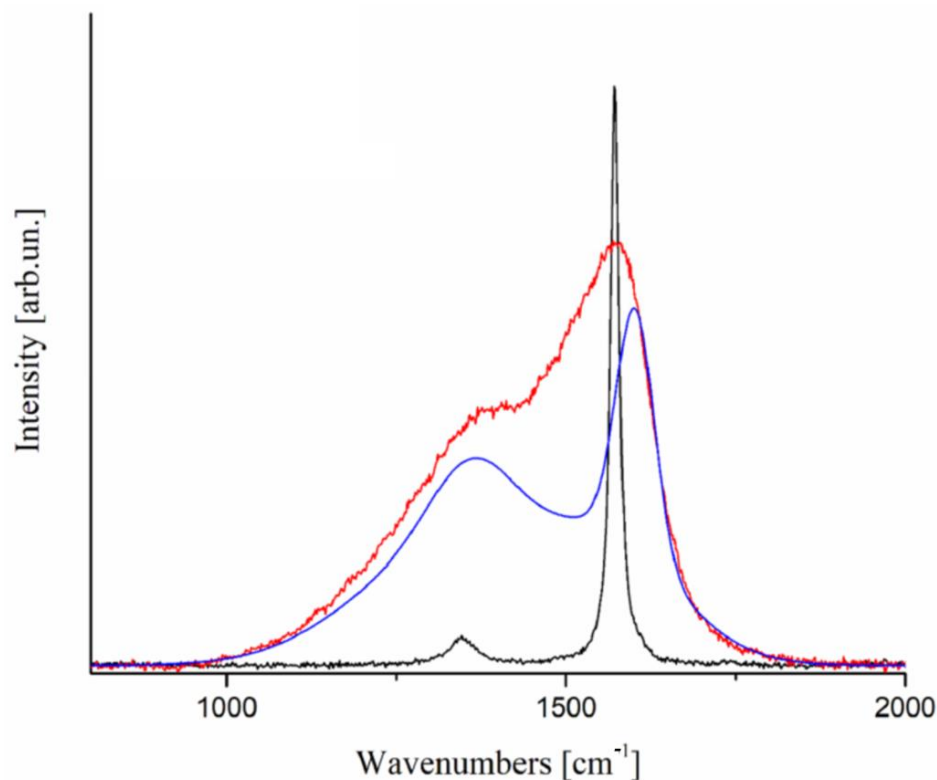


Figure 5. Raman spectra of graphite (black), disordered pyrolytic carbon (blue), and diamond-like carbon (red), as reported by Tagliaferro et al. [118] (CC BY 4.0).

The evolution of G and D peaks could be used to investigate the temperature used for the pyrolytic conversion of polymeric feedstock [119–124], but also to evaluate the different feedstock used for the production of carbon [125]. The main issue related to the interpretation of pyrolytic disordered carbon is represented by the fitting procedure that is affected by the uncertainty in the number and the lineshape of the components used. Tagliaferro et al. [118] introduced a symmetric lineshape named GauLor, composed of a central domain with a Lorentzian lineshape linked with to Gaussian tails. GauLor lineshape perfectly matches the stretched exponential decay and represents the first physical approach to overcome the arbitrary use of any kind of lineshape.

3.2. Inorganic Materials

3.2.1. Biological Applications

Metals, ceramics, and polymer nanoparticles may present interesting properties for biological applications such as biocompatibility [126], magnetism [127], and photoluminescence [128]. Their characterization with Raman spectroscopy can reveal important features of the materials under study, for example by probing the interaction between the nanoparticles and the biological environment.

Bhaumik et al. [129] successfully investigated the nano–bio boundaries between ZnO nanoparticles and adenosine triphosphate (ATP; the major player in the energetic

transport in cells), showing a pH-dependence on the interaction. Zinc oxide nanoparticles are nontoxic, biocompatible, and usable in drug delivery, medical materials, and cosmetics. Furthermore, the Raman activity of ZnO is very high, making this technique a good probe for studying their interaction with biomolecules. Considerable Raman shifts related to the interaction between ZnO and NH₂ bonds in the adenine ring, N7 atom in the adenine ring, and the phosphate bonds have been detected in acidic environments.

Agressott et al. [130] studied protein-coated silver nanoparticles synthesized by common fungi (*Rhodotorula mucilaginosa* and *Rhodotorula glutinis*) following a green and biocompatible approach. Using TERS, they successfully investigated the size and the morphology of the protein capping layer, how it interacts with the Ag nanoparticles (the sulphur bridge in amino acids was detected), and how the whole system interacts with the visible light.

The detection of uric acid using silver nanoparticle-coated ZnO/Fe₃O₄ composites was performed by Alula et al. [131] using surface enhanced Raman spectroscopy (SERS). A linear response was obtained, for up to 10 μM concentrations, for an aqueous solution of uric acid. Finally, an application to determine the uric acid concentration in a real sample of non-treated urine was performed.

Caprara et al. [132] studied the DNA molecule formation using SERS with single DNA chain coated gold nanoparticles, comparing the Raman spectra of the final solution with the starting one. This paved the way for a bottom-up approach for the design of tailored nanostructures, taking advantage of the programmable assembly of DNA. The SERS response showed the successful formation of ordered nanostructures with a controlled interparticle distance.

An interesting improvement with respect to TERS is represented by the SHINERS technique, as reported by Li et al. [59]. It is based on the covering of a targeted substrate with silica encapsulated noble metal nanoparticles. This approach could be combined with portable instrumentation to boost the on-field analysis sensitivity. Furthermore, it has been used by El-Said et al. to detect two of the main neurotransmitters in the human brain (γ-aminobutyric acid and glutamate) that are important in the diagnosis of neurological diseases, using gold nanobipyramids coated with a thin layer of polypyrrole [133]. This process can be used to probe concentrations of γ-aminobutyric acid and glutamate in human serum within a wide range of concentrations.

Erythrosine B coated gold nanoparticles were used by Pinilla-Peñalver et al. [134] as a probe to sense both compounds as a result of the high Raman activity of the dye. For the Au NPs, a linear response was obtained in the 1–12 ng/L range, while for erythrosine B, the linear region resulted in between 5 and 150 μg/L. The detection limits were at 0.3 ng/L and 1.4 μg/L, respectively.

Raman spectroscopy can also be used for the direct sensing of biomolecules such as DNA, proteins, lipids, and carbohydrates. Kuhar et al. [135] showed that this technique can successfully probe inside a set of pure biomolecules: bovine serum albumin (BSA), calf thymus DNA, cholesterol, and glucose. Several unique marker bands have been identified and reproduced.

Butler et al. [136] proposed a detailed protocol to perform Raman spectroscopy on biological samples (i.e., plant tissues, cells, biofluids, etc.), both in vivo and in vitro. The results obtained included a mapping of a portion of endometrial tissue and a classification of blood plasma and serum.

The diagnosis accuracy of tumors can be significantly improved using Raman spectroscopy [137]. Evidence was found for the diagnosis of enchondroma and chondrosarcomas [138], thyroid cancer [139], lung cancer [140], ovarian cancer [141], breast cancer [142], colorectal cancer [143], prostate cancer [144], and many others [145]. The advantage of this technique is that the analyses are carried out directly on biological samples that are often easy to obtain (serum, blood, saliva, urine, sperm, etc.), and the results are obtained in a short time. Furthermore, several pathologies and infections can be detected through this technique, and this is especially important in a pandemic event. An example is offered by

the research of Ryzhikova et al. [146], who describe a way to anticipate the diagnosis of Alzheimer's disease using Raman spectroscopy. This is important because this pathology can nowadays be diagnosed through specific neuropsychological tests, but only in the advanced stage. The development of the technique cited in this work could reduce diagnosis times and anticipate treatments and prevention.

Nonalcoholic fatty liver disease is a chronic disease that occurs in two stages. The first phase (NAFL) consists of an initial accumulation of fat in the liver, the second phase (NASH) is steatohepatitis, a degenerative form that can lead to liver cirrhosis and cancer. Gurian et al. [147] developed a SERS-based diagnostic system to discriminate between NAFL and NASH by testing 5 μ L of blood plasma on an Ag nanoparticles substrate. The short response time and the ability to differentiate indicates to researchers that SERS could be an excellent method for future diagnoses of NAFL and NASH.

Active tuberculosis and latent tuberculous infection can be detected with both Raman spectroscopy and SERS. Kaewseekhao et al. [148] developed a protocol based on Raman spectroscopy to diagnose and distinguish between active and latent tuberculosis with an accuracy of 91.15%. Furthermore, SERS can distinguish between all TB-infection categories with 100% sensitivity and specificity.

Raman spectroscopy of urine samples can diagnose chronic renal failure (CRF) disease with an acceptable accuracy, as shown in the work of Chen et al. [149], in which genetic algorithms aided in RS obtained accuracies of between approximately 65% and 85%.

Hepatitis C virus (HCV) was detected in blood serum samples from positive patients and healthy individuals using Ag nanoparticles as a substrate for SERS. A similar study showed that Raman spectroscopy is not able to identify single RNA groups of the virus, while SERS is able to [150]. Furthermore, surface enhanced Raman spectroscopy can detect different levels of viral flow in the analyzed biological cells with a good sensitivity.

The Hepatitis B virus has a distinctive Raman signature, and can be detected with SERS by developing algorithms for data processing and recognition, achieving a 98.82% accuracy [151].

Picomolar concentrations in the biofluids of biomarkers of traumatic brain injuries were measured by Rickard et al. [152]. Sub-micrometric pillars covered with a gold nanolayer made it possible to achieve a considerable sensitivity using a hand-held, optofluidic SERS device.

Cardiac troponins are isoforms released by the heart into the blood, especially when it has been damaged. Their concentration in the blood is proportional to the extent of the damage suffered. A troponin LOD has been demonstrated between 7.6×10^{-4} ng/L and 800 ng/L for SERS devices, whereas the available LOD of laboratory and point-of-care testing for cardiac troponin ranges from 100 to 800 ng/L. The device of Su et al. [153] used a three-dimensional ordered macroporous coupling Au–Ag–Au plasmonic array substrate in addition to Ag–Au nanostars as nanotags [154].

Biomarkers of amyotrophic lateral sclerosis (ALS) could be detected in the saliva of patients using Raman spectroscopy [155]. Such an approach can drastically shorten diagnosis times, which are lengthened by symptom checks, which are often very similar between different pathologies (Alzheimer's disease, Parkinson's disease, ALS, etc.). Thus, a fingerprint of the different biomarkers could lead to precise and quick diagnoses of the most dangerous neurodegenerative diseases.

Recently, Acri et al. [156] proposed a method to diagnose the pediatric onset of inflammatory bowel diseases, such as ulcerative colitis and Crohn's disease. Their approach involved taking fecal samples to be analyzed with Raman spectroscopy and compared with the fingerprints of the various pathologies. The secondary structure of some proteins is modified by inflammation—there is a greater ratio between non-reducible and reducible cross-linking in the case of infection, which can be detected by Raman spectroscopy.

The COVID-19 pandemic has changed the lives of billions of people around the world. One of the challenges is to develop a quick, effective, and low-cost test to diagnose this virus as well as those to come. Jadhav et al. [157] proposed a microfluidic device coupled with

Raman spectroscopy to trap the virus collected by biological secretions (saliva, tears, and nasopharyngeal fluid) and study its SERS signature. The signal enhancement is produced by Ag/Au nanoparticles that cover the carbon nanotubes placed in the microfluidic device.

Gang Yin et al. [158] proposed a Raman approach to identify Coronavirus infection in human serum. A machine learning support-vector machine (SVM) was used to create a fingerprint database, in order to improve the quality and speed of diagnosis over time.

Another approach was recently proposed by Elsharif et al. [159], using a layered surface to detect SARS-CoV-2 in a biological sample. The SERS structure is composed of a polydimethylsiloxane nanoarray creating a nanocavity in which light reflects on the sample several times. A gold nanorod substrate is then placed on the top, which could improve the SERS influence by three orders of magnitude.

A very interesting field of study is represented by the Raman spectroscopy of a living cell [160]. As reported by Li et al. [161], this approach could be used to evaluate the functionalities of lipid droplets and bacteria by analyzing the finger print region showing the relative abundance of chemical functions, as reported in Figure 6.

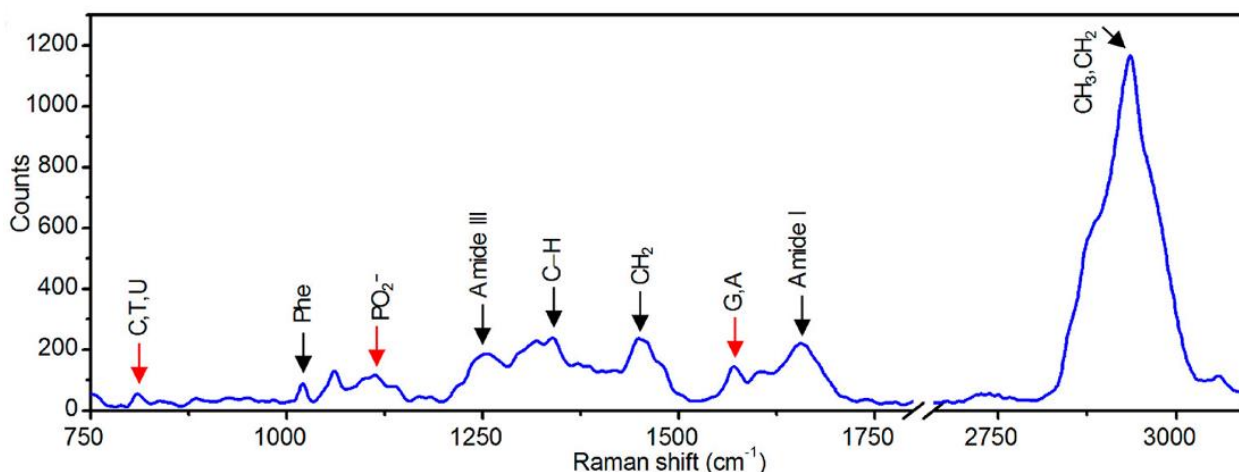


Figure 6. Raman spectrum of *S. aureus* cells in a panel acquired by using an excitation power on the sample of up to 2 mW and an acquisition time of up to 30 s, as reported by Li et al. [161]. Picture reproduced with permission from the American Chemical Society (Copyright © 2019, American Chemical Society).

This approach was also used to monitor the organelle in vivo [162,163], the chromosomes [164], and for the evaluation of the hemoglobin in red blood cells and for the density of water inside the cells [165]. This last study is particularly interesting because it proved the presence of a strong and highly organized hydrogen bond network in the cytosol.

Raman spectroscopy has also been used for analyzing the structural features of epithelial tissue [166,167] in order to detect and diagnose precancerous changes [168].

3.2.2. Mineralogy

Raman spectroscopy, being a non-destructive, precise, and fast technique, is a common tool in mineralogy for the analysis of a large variety of samples. Furthermore, the Raman analysis of a mineralogical sample can provide a punctual, linear, surface, and in some cases even three-dimensional mapping of the analyzed sample. Due to these characteristics, Raman spectroscopy is used to analyze jewels, minerals, works of art, and historical artifacts.

Nasdala et al. [169] reported that this technique can investigate inside the inner structure of mineralogic samples in different temperature and pressure conditions. These in situ experiments can be interesting in the case of a sample with temperature and pressure related properties. Raman band shifts can be used for indirect measurement of the condition of the sample under analysis. This led Schmidt et al. [170] to present a Raman-based pressure sensor that can work even at high temperatures, at which the ruby luminescence

technique does not. Their research discovered a linear dependence of the ν_3 Raman band of fully crystalline synthetic zircon ($Zr_{0.987}Hf_{0.013}[SiO_4]$) with pressure between 0.1 MPa and 6.6 GPa.

Raman spectroscopy also helps rovers on other planets to analyze extraterrestrial rocks. The Perseverance rover is equipped with two Raman instruments: SHERLOC (Scanning Habitable Environments with Raman and Luminescence for Organics and Chemicals) and SuperCam. The first is located on the arm turret of the rover and is associated with an X-ray fluorescence instrument. It works at an excitation wavelength of 248.6 nm (deep UV) and is used to study the surface of the samples taken. The second includes laser-induced breakdown spectroscopy (LIBS), time-resolved Raman and luminescence spectroscopy, and visible–infrared (VISIR) spectroscopy [171].

Yiheng et al. [172] scanned the lunar meteorite NWA 10480, found in 2015, with Raman, finding shifts related to olivine, plagioclase, pyroxene, and Fe-Ti oxide. The data elaborated from the 650+ collected 2D spectra showed that NWA 10480 consists of 49 vol.% pyroxene, 20 vol.% plagioclase, 15 vol.% olivine, and 3 vol.% Fe-Ti oxides.

A similar work was done by Haijun et al. [173], who performed a Raman analysis on the NWA 4884 meteorite, reporting a large amount of Raman modes for the mineral phases of the sample. The composition of the analyte appeared to be very different from that of the study mentioned above. In fact, the results showed a composition of 42.4 vol.% pyroxene, 26.6 vol.% plagioclase, 20.7 vol.% olivine, and only traces of Fe-Ti oxides. These results are an average taken along the whole sample, which consisted of domains in which a single phase prevailed among the others. This proved the great lunar mineralogical variety and that the lunar meteorites arriving on the Earth's surface may come from different and distant areas of our satellite.

Polavaram et al. [174] proposed an innovative protocol to create 2D maps of mineralogical samples. The method proposed in this work greatly reduce the analysis times as it is able to analyze non-polished samples with an auto-focusing Z-mapping feature that allows for identifying different polymorphs with great accuracy (<97%) and a high spatial resolution (<0.3–2 μm).

Raman spectroscopy has been used as new certification method for precious minerals (especially diamonds, rubies, and sapphires) [175]. The idea consists of marking the mineral using a femtosecond laser, and creating a QR code in which all the information on the material (carat weight, origin, processing, etc.) is saved. Raman mappings and scanning electron microscope images revealed that the dots had a diameter of 14 microns and were spaced 14–18 microns apart. The effectiveness of the marking can be seen from the Raman spectra. For example, diamond exhibits graphitization in the areas exposed to the laser and this is very noticeable from the Raman spectrum. In fact, here, the classic Raman peak of the diamond (sp^3 carbon) located at about 1340 cm^{-1} is accompanied by the D and G bands, which underline the presence of the sp^2 bond. This is further confirmed by the presence of second order peaks at 2740 cm^{-1} (2D peak) and 2940 cm^{-1} (D and G peak).

Culka et al. [176] proposed an easy, fast, and portable method for the in situ identification of minerals using a portable sequentially shifted excitation Raman spectrometer and the online RRUFF database. The Raman apparatus automatically removes the fluorescence background and compares the obtained spectra with the RRUFF database for matching minerals. In the test proposed by the researchers, the system recognized 19 out of 20 mineral species, confirming itself as a good method for the in situ identification of gemstones.

3.3. Advanced Applications

3.3.1. Cultural Heritage

Raman spectroscopy has gained great attention in the field of cultural heritage diagnostics [177]. The non-destructive Raman analysis has been used to evaluate the ageing and degradation of plenty of precious artifacts, paints, and statues [178].

Prieto et al. [179] investigated, through Raman spectroscopy, a relatively recent paper-based artifact, a 18 century Spanish playing card found in Perú. The authors were able

to identify the composition of the inorganic pigments used and the degradation of the cellulosic materials. Cellulose-based historical artifacts represent a very interesting family for Raman investigation [180,181]. The analysis of pigments used in medieval manuscripts enlightens modern researchers about the techniques developed in the monastery during the middle ages [182,183]. Micro Raman analysis could be also used to explore the ageing of far older cellulosic materials, such as a three-thousand-year-old mummy dating back to the late period of ancient Egypt. Mummies are another relic of the past that have been deeply analyzed using Raman spectroscopy. Several studies have investigated organic tissue ageing in different and far regions of the global [184,185]. Ancient inorganic-based artifacts represent another field of interest for Raman analysis. Barone et al. [186] investigated the composition of a Roman jewels collection, identifying each gemstone used.

Colomban et al. [187] investigated the arsenic content in European porcelain. By using Raman spectroscopy, the authors were able to build a timeline of the evolution of ceramic pigments across centuries in the European region.

The Raman analysis of painting is very interesting considering the amount of information recovered from a simple Raman spectrum.

Antunes and co-workers [188] were able to evaluate even the ground layers behind 15–16 century Portuguese paint through micro Raman analysis. The authors identified gypsum and inorganic anhydrite with a good match. Furthermore, wall painting could be analyzed as reported by Appolonia et al. [189]. The authors described with a great accuracy the inorganic species used in the realization of wall paint preserved in the northwest of Italy, finding a good agreement between Raman outputs and X-ray fluorescence investigation.

Raman spectroscopy has also been used for evaluating bronze patinas, reporting the negative effect of several environmental conditions [190,191]. Based on this study, a more efficient preservation policy could be enforced to limit the environment degradation of rare and ancient artifacts.

3.3.2. Quality Control and Analytical Processes

Raman spectroscopy could also be used as a quality control tool for monitoring chemicals crystallization [192]. Roozeboom et al. [193] reported a study about the use of the Raman technique as an analytical tool for monitoring the precipitation and crystallization of different types of zeolite. The authors described the disappearance of complex aluminum hydroxide with the advancement of zeolite formation, and they were able to discriminate between zeolite A, B, and Y by observing the signals in the region from 300 cm^{-1} up to 680 cm^{-1} . The production of pharmaceuticals could be also monitored by the Raman technique, as reported by Strachan et al. [194]. In this field, the ability of Raman spectroscopy to identify the different crystalline configurations of bio-active molecules represents a solid point in favor of the use of this approach [195]. Wang et al. [196] estimated the shelf life of aspirin tablets through simple Raman analysis, achieving results comparable with those from high performance liquid chromatography. As a quality control tool, Raman spectroscopy was used by Hali et al. [197] to evaluate the distribution of paracetamol and caffeine in drug grains. The authors were able to provide a positive/negative control regarding the presence of the correct ratio between the active bio-molecule identifying the unacceptable materials.

Raman spectrometers are also used in the food industry as rapid on-line and in-situ quality control for fish, meat, and drink production [198].

Kneipp et al. [199] firstly reported single molecule SERS that offers significant advantages when compared with single molecule fluorescence [44]. SERS is also developed for use in spectroelectrochemistry, and in conjunction with electrochemistry, it allows for detecting the behavior of molecules in different oxidation states. The structural changes and redox properties of the molecule tetrathiafulvalene were monitored by SERS [200]. Dynamic SERS imaging inside a living cell can be examined by means of a gold nanoparticle, which travels through the intracellular space to probe the local molecular information. Simultaneous tracking of the particle motion and SERS spectroscopy provides molecular

maps of organelle transport and lysosomal accumulation [201]. SERS imaging can provide chemical information with a spatial resolution in the micrometer range, and label-free imaging of biofilms has also been performed [202]. In addition, the SERS capability to detect the chemical vapor for the sensing of highly toxic molecules, such as chemical warfare agents and toxic industrial chemicals, has been demonstrated. Overcoming the lack of interaction between of these molecules with the surface, detection was run using a device such as a combined microfluidics-SERS sensor [203]. Explosives such as the half-mustard agent [204] and dinitrobenzenethiol [205] have been successfully detected using SERS. Of significant interest is the transition of SERS detection from the lab to the field [206]. The advent of portable Raman spectrometers allows for on-field SERS measurements of real-time environmental chemical analysis and monitoring, ranging from the structural characterization of soils, through ultrasensitive detection of pollutants and heavy metal ions, to the analysis of plants, tissues, and microorganisms [207]. Stand-off detection by SERS is made possible with optical fiber probes [208], relevant for in vivo measurement or biomedical applications involving offset measurements through the skin, and the stand-off molecular detection from samples at distances of 15 m [206]. The detection and identification of dilute bacterial samples by SERS has been explored by mixing aqueous suspensions of bacteria with a suspension of nanocolloidal silver particles. The detection limit of this technique is $\sim 10^3$ cfu/mL, which is quite promising compared with conventional detection methods [209].

SERS is also a promising method for the ultrasensitive detection of chemical species that are relevant to homeland security [210].

4. Considerations about the Limitations and Advantages of Raman Spectroscopy

Raman spectroscopy has found many applications in advanced materials science characterization. Nonetheless, its spread has led to a several misunderstandings in data elaboration and interpretations. The best example of this misleading approach is represented by the elaboration of disordered carbon. In the scientific literature, many papers faced the fitting of a Raman signal using an arbitrary number of components [211–213], selecting lineshapes without a solid physical reason [214,215].

A more general and neglected issue is represented by the baseline correction of the spectra. Baseline is the first and most crucial step for any Raman spectra elaboration, and even though many approaches have been developed [216–219], not enough researchers use them in an appropriate way.

Moreover, Raman quantitative analysis may be unreliable if not compared with and validated by supported well established procedures [220].

Nonetheless, Raman spectroscopy could provide a very good insight for investigating many materials, and with the creation of large databases, qualitative Raman analysis will be more and more reliable [221–224].

A concise summary of the information and limitations encountered for several materials and stressed in manuscript are reported in Table 1.

Table 1. Overview of the data achievable for each of the main classes of materials analyzed through Raman spectroscopy.

Materials	Features	Informations Provided	Limitations
CNTs	<ul style="list-style-type: none"> ■ D and G peaks ■ RBM ■ 2D region 	<ul style="list-style-type: none"> ■ Order (I_D/I_G) ■ Chirality (ω_{RBM}) ■ Metallic or semiconductive behavior (G peak) ■ Number of layer (G peak shape) 	<ul style="list-style-type: none"> ■ MWCT signal fits could be arbitrary in the choice of lineshape ■ RBM detection in MWCNT is possible only for straight isolated CNTs on a SERS substrate

Table 1. Cont.

Materials	Features	Informations Provided	Limitations
Graphene and graphene like materials	<ul style="list-style-type: none"> ■ G peak ■ D peak (only for defective structures) ■ 2D region 	<ul style="list-style-type: none"> ■ Defects (rise of D peak) ■ Order (I_D/I_G) ■ Number of graphene layers (D and G' peaks) 	<ul style="list-style-type: none"> ■ Difficult discrimination between few-layer defected graphene and good quality graphite
Amorphous carbon	<ul style="list-style-type: none"> ■ D and G peaks ■ 2D region 	<ul style="list-style-type: none"> ■ Order (I_D/I_G), shape of region 2-D ■ Analysis of reorganization based on production temperature 	<ul style="list-style-type: none"> ■ Fit of D and G peaks ■ Physical explanation of D and G components
Inorganic species	<ul style="list-style-type: none"> ■ Several peaks at low Raman shifts (generally lower than 1000 cm^{-1}) 	<ul style="list-style-type: none"> ■ Chemical species identification ■ Phase identification 	<ul style="list-style-type: none"> ■ Difficult spectra processability ■ Need of supporting informations for trustworthy reference databases (i.e., XRD)
Organic molecules	<ul style="list-style-type: none"> ■ Several peaks based on the specie analyzed 	<ul style="list-style-type: none"> ■ Structural informations of macromolecule (molecular weight, ageing, and crystallinity) ■ Crystallization monitoring ■ Quantification even on molecular level (SERS) 	<ul style="list-style-type: none"> ■ A peak-by-peak assignment is very challenging for macromolecule ■ Very sensible to interferences of non-targeted organic species ■ Quantification through SERS techniques required and external validation (i.e., chromatography and quantitative spectroscopy)

5. Conclusions

Raman spectroscopy is a very powerful and polyhedral tool that could be used for the investigation of many fields and many more materials.

As we discussed, Raman is useful for advanced materials science and real-life applications, always providing a deep and meaningful insight into the properties and arrangement of the sample analyzed. Considering the decreasing cost of Raman instruments and their portability, the spread of on-field Raman analysis will mark a new game changing event establishing the advanced materials characterization as a routine analysis.

Considering the vast literature discussed above, we firmly hope that this review will represent a reference point for newcomers and a useful tool for field experts as well.

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